LECTURE 2: Oxide ion conductors

Functional Inorganic Materials
Fall 2020

Tuesdays: 10.15 - 12.00
Thursdays: 10.15 - 12.00
Remote Zoom lectures

#	Date	Who	Topic
1	Tue 27.10.	Maarit	Introduction + Superconductivity: High-T_c superconducting Cu oxides
2	**Thu 29.10.**	**Maarit**	**Ionic conductivity (Oxygen): SOFC & Oxygen storage**
3	Tue 03.11.	Maarit	Ionic conductivity (Lithium): Li-ion battery
4	Thu 05.11.	Maarit	Hybrid materials
5	Tue 10.11.	Antti	Thermal conductivity
6	Thu 12.11.	Antti	Thermoelectricity
7	Tue 17.11.	Antti	Ferro-, pyro-, and piezoelectricity
8	Thu 19.11.	Antti	Magnetic and multiferroic oxides
9	Tue 24.11.	Mady	Metal-based energy-saving applications
10	Thu 26.11.	Mady	Metal-based energy-efficient windows and solar absorbers
11	Tue 01.12.	Mady	Metal oxides for energy-saving applications: Past and new trends
12	Thu 03.12.	Mady	Materials design and new perspectives

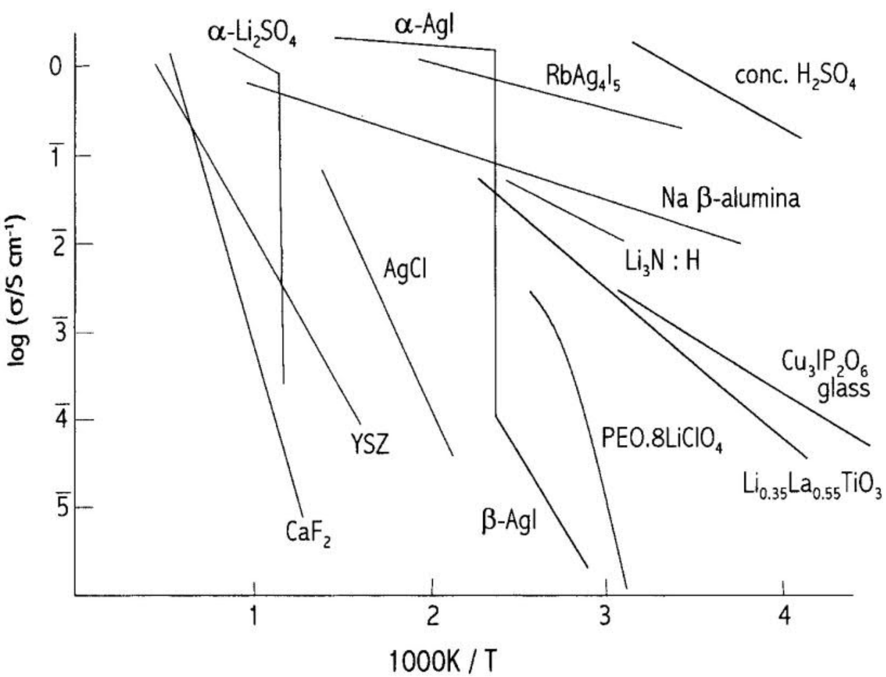
#	DATE	TOPIC & KEYWORDS
1	**27.10.**	**(High-T_c) Superconductivity** New-material design, Multi-layered crystal structure, Mixed-valency, Oxygen nonstoichiometry
2	**29.11.**	**Ionic conductivity: Oxygen** Oxygen vacancies, Redox-active cations, Mixed valency, Cation substitutions (isovalent/aliovalent), Crystal symmetry, Oxygen storage, SOFC
3	**03.11.**	**Ionic conductivity: Hydrogen & Lithium** Water absorption & Oxide/hydroxide, Li-ion battery, Solid-state electrolytes
4	**05.11.**	**Hybrid materials** Inorganic-organic materials, MOF, ALD/MLD, Layer-engineering

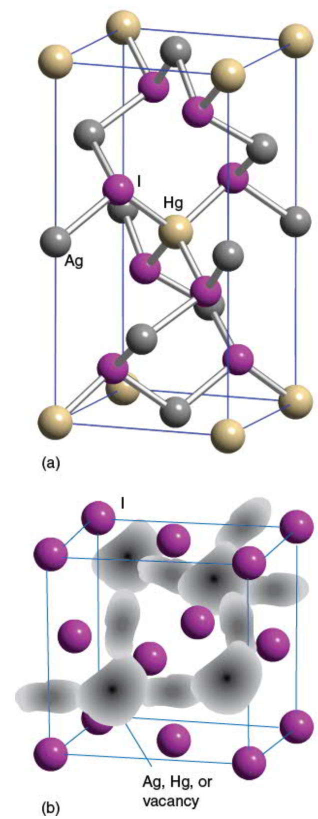
LECTURE EXERCISE 2

1. (a) Are all oxide-ion conductors good oxygen-storage materials ? Justify !
   (b) Are all oxygen-storage materials good oxide-ion conductors ? Justify !
2. Explain the differences of the two compounds, $YBa_2Cu_3O_{7-\delta}$ and $YBaCo_4O_{7+\delta}$, regarding the oxygen non-stoichiometry, i.e. how they absorb/desorb oxygen upon heating.
3. You can use Zr, Ce or Y oxides or their mixed compounds. How would you construct your material if you need to make (a) good electrical conductor but bad ionic conductor, (b) good ionic conductor but bad electrical conductor, (c) good electrical and good ionic conductor; explain shortly why.
4. What are the requirements (qualitatively) in terms of electrical conductivity and ionic conductivity for SOFC (i) cathode, (ii) anode, and (iii) electrolyte ?
5. You have four (hypothetical) perovskite compounds, $La(Ga,Mg)O_{3-\delta}$, $(La,Y)GaO_{3-\delta}$, $(La,Sr)GaO_{3-\delta}$ and $(La,Sr)CrO_{3-\delta}$. Explain and justify (for all the four materials separately) whether the material could show some promise as a SOFC cathode or electrolyte material.

IONIC CONDUCTIVITY

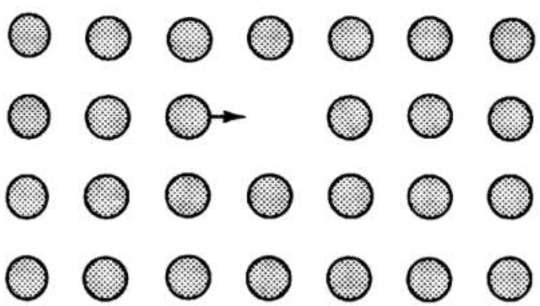
- Movement of ions in **solid** (or liquid): defective/disordered crystals, glasses, polymers, nanocomposites, gels
- Ion conductivity increases with increasing temperature
- Faraday 1839: laws of electrolysis apply to ionic solids: e.g. PbF_2 & Ag_2S
- Prototype superionic conductor: high-temperature (>147 °C) disordered AgI
- Ford Motor Co.: BASE (β-alumina solid electrolyte): Na-Al_2O_3 → Na-S battery
- Present examples: Ag_2HgI_4 (Ag^+-ion), LaF_3 (F^--ion)



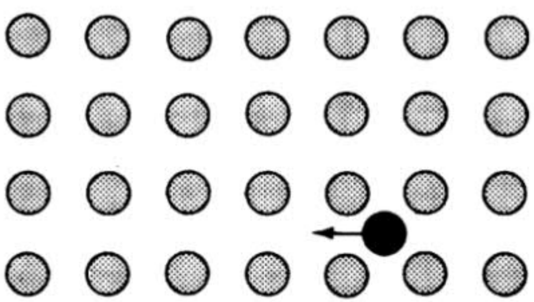


IONIC CONDUCTOR (solid)

- <u>Other terminologies</u>:
  Fast ion conductor, Superionic conductor, Solid electrolyte, Solid state ionics
- Highly mobile ions move/hop through an otherwise rigid crystal structure
- Fast ion conductors are intermediate between ***regular crystalline solids with immobile ions*** & ***liquid electrolytes*** (without a regular structure) ***with fully mobile ions***
- Measurement of ionic conductivity: electrochemical impedance spectroscopy (EIS)
- **<u>APPLICATIONS</u>:**
  batteries, fuel cells (e.g. **SOFC**), supercapacitors, chemical sensors, separation membranes, gas (e.g. **oxygen) storage**



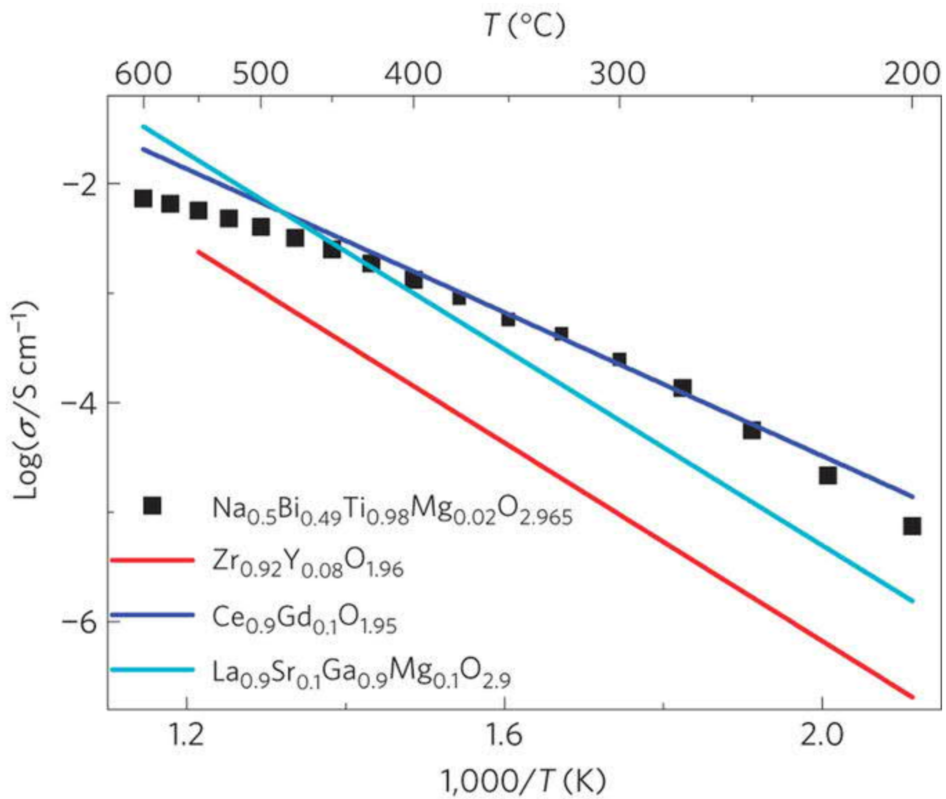
mobile vacancy



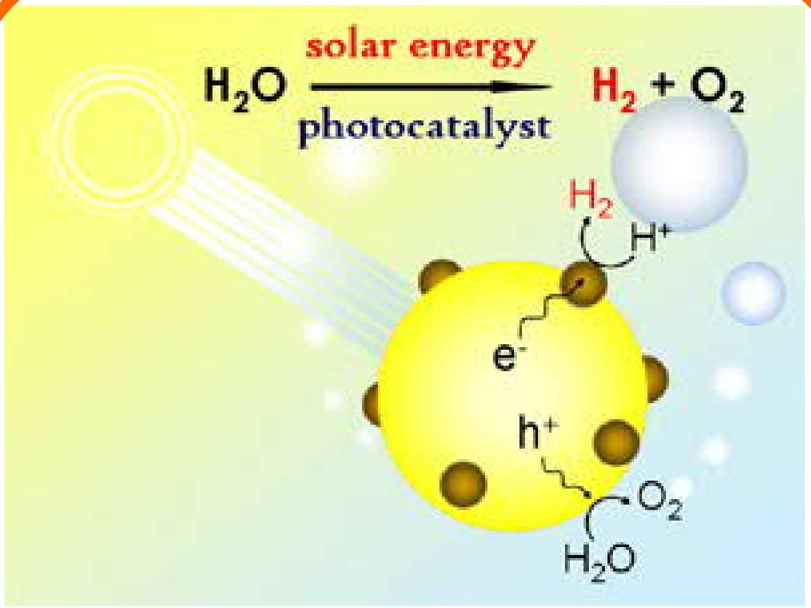
mobile interstitial

<u>EFFECTIVE IONIC RADII</u>	
O^{2-}	**140 pm**
OH^{-}	**137 pm**
Li^{+}	**60 pm**
H^{+}	**very small**

	Material	Conductivity (S m⁻¹)
Ionic conductors	Ionic crystals	$< 10^{-16} – 10^{-2}$
	Solid Electrolytes	10^{-1}-10^{3}
	Liquid electrolytes	10^{-1}-10^{3}
Electronic conductors	Metals	10^{3}-10^{7}
	Semiconductors	10^{-3}-10^{4}
	Insulators	$< 10^{-10}$



Solid oxide fuel cell (SOFC)
oxide-ion conducting materials
NISSAN Motor Co. Ltd.



H_2/O_2 separation
in photocatalytic water splitting
Domen Lab. (Univ. of Tokyo)

Redox exhaust gas catalyst
TOKYO ROKI Co. Ltd.

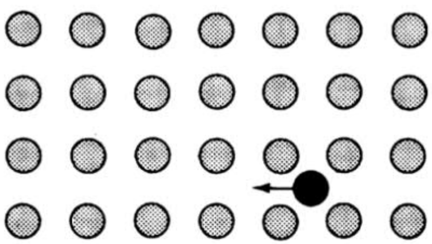
- Separation
- Purification
- Sensors

Examples of APPLICATIONS

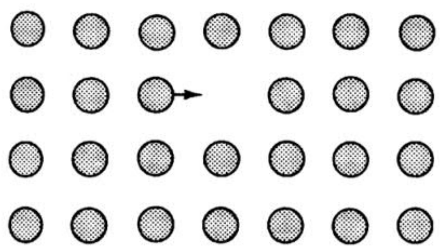
OXIDE-ION CONDUCTIVITY & OXYGEN-STORAGE

MATERIAL DESIGN CONSIDERATIONS

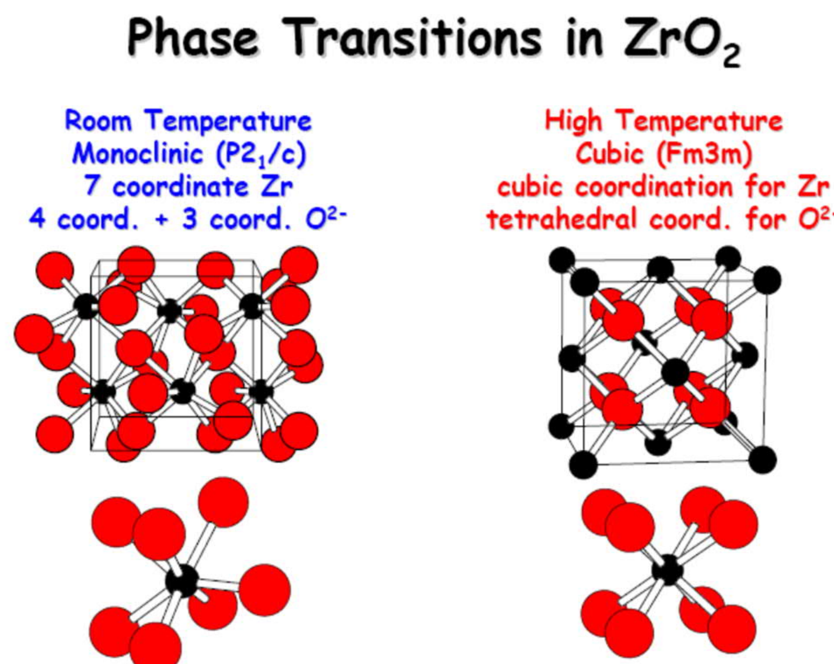
- **Open structure**: space for O^{2-} ion diffusion
- **Oxygen vacancies**: efficient O^{2-}-ion hopping
- **High crystal symmetry**: all oxygen sites equivalent (e.g. cubic ZrO_2 desired)
- **Redox-ative cations**: oxygen-content variation → **OXYGEN STORAGE**
- **For sensor and separation applications → thin films**
- **Other important factors: chemical/thermal stability, thermal expansion, …**

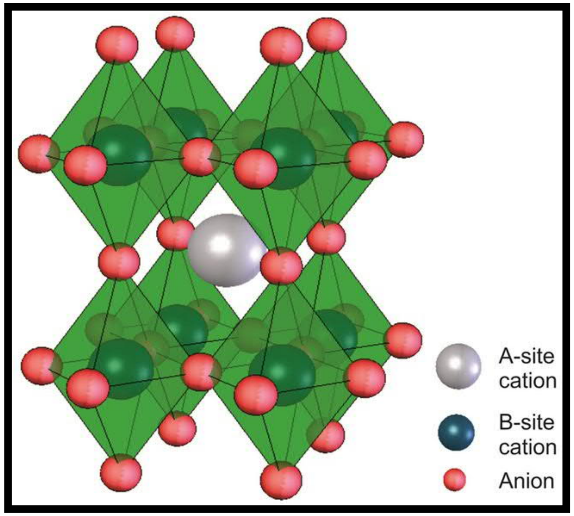


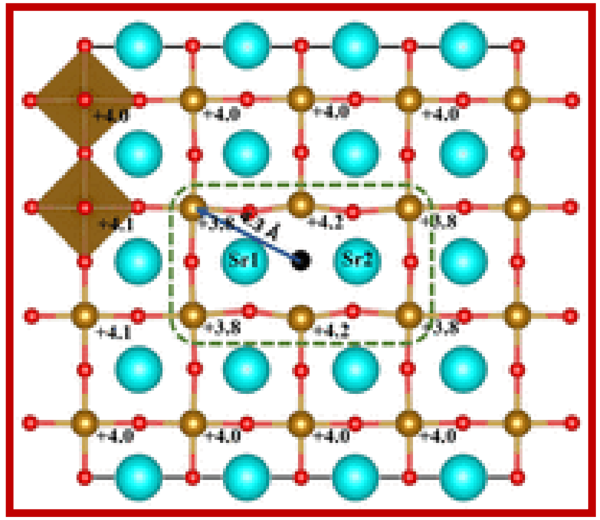
mobile interstitial



mobile vacancy

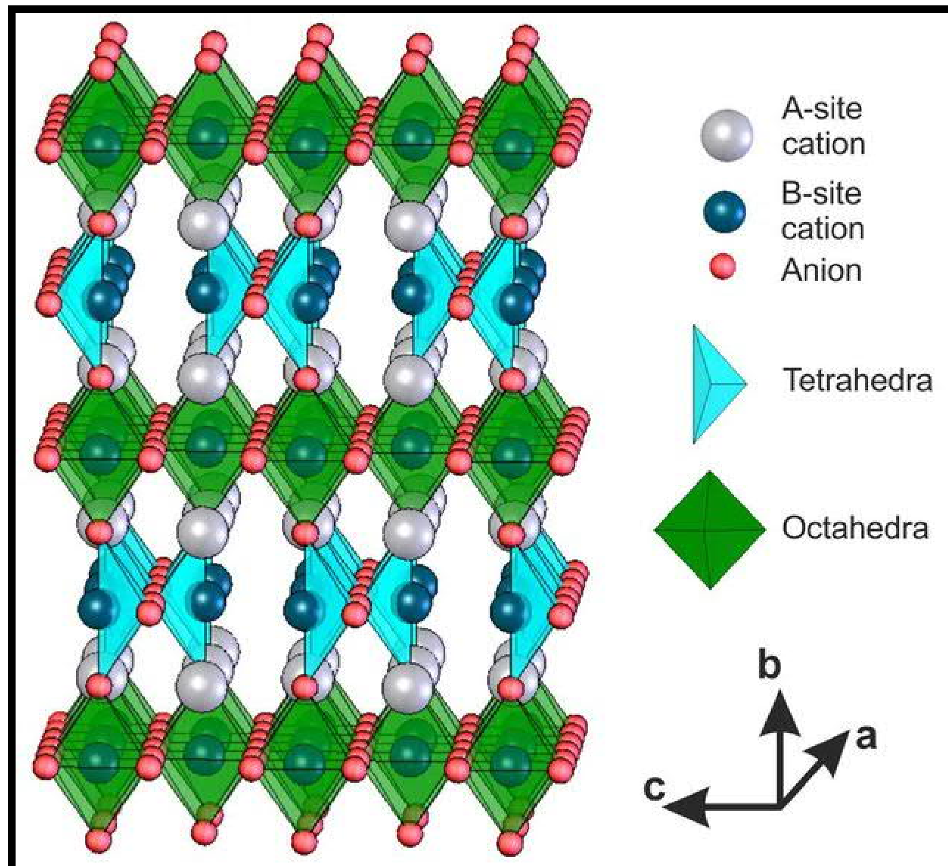






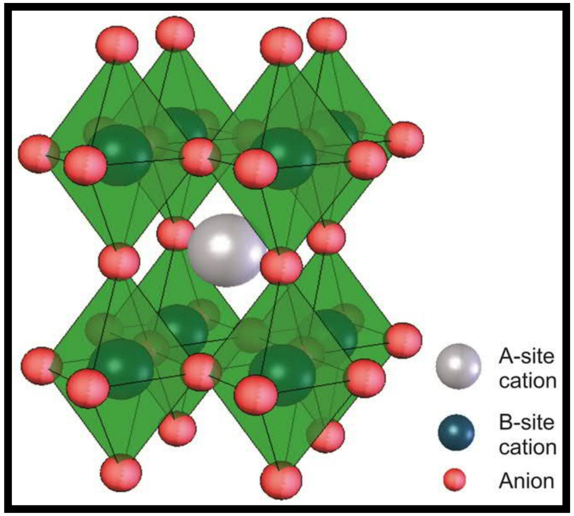
Perovskite $ABO_{3-\delta}$

- Prone for oxygen vacancies



Oxygen-vacancies ordered:

- Brownmillerite $ABO_{2.5}$

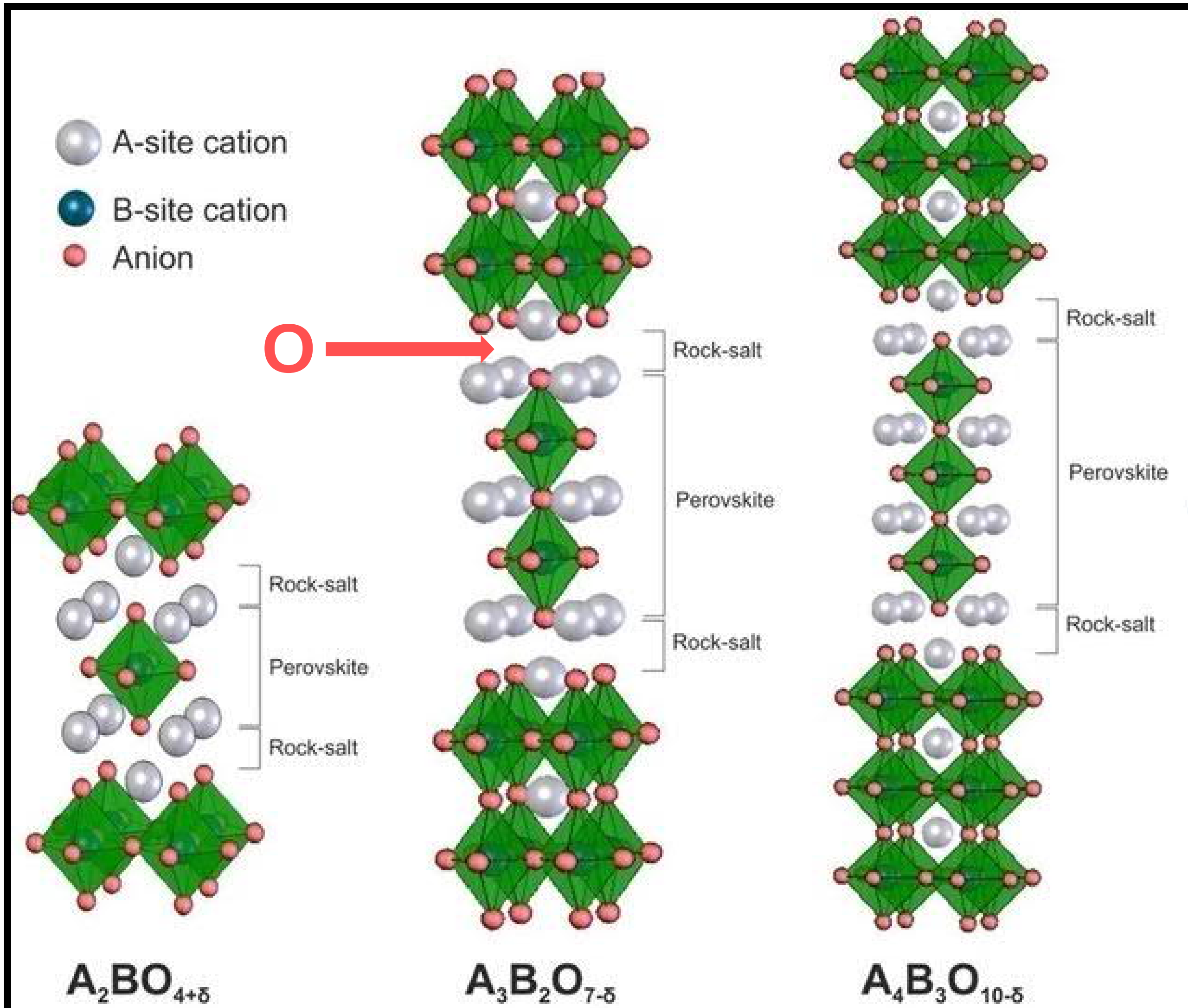


Perovskite ABO_3

- No space for interstitial oxygen

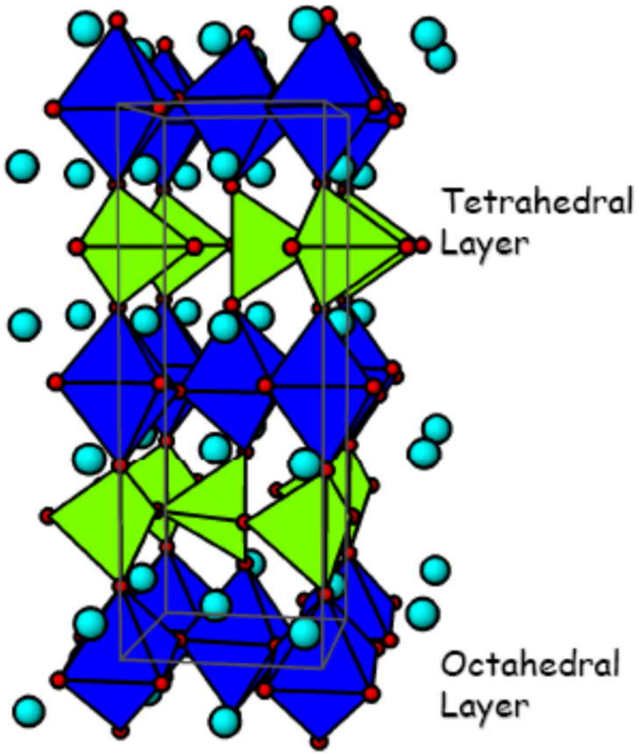
Ruddlesden-Popper $A_{n+1}B_nO_{1+3n}$

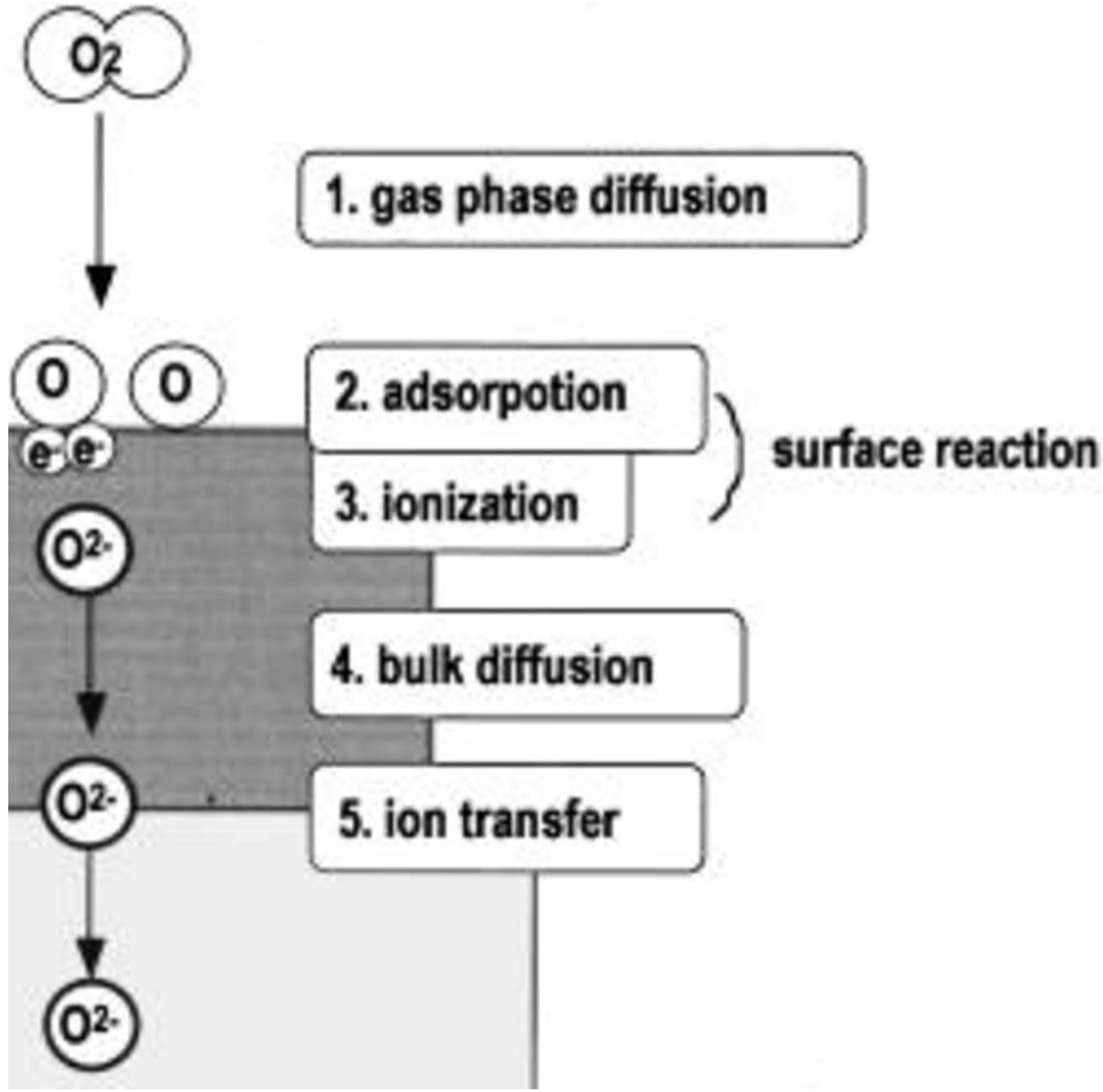
- Enough space for interstitial oxygen



$Ba_2In_2O_5$ ($BaInO_{2.5}$)

- **Brownmillerite structure**
  derived from the perovskite structure (oxygen vacancies ordered into layers)
- **At 800 °C oxygen vacancies disorder and the ionic conductivity jumps from 10^{-3} S/cm to 10^{-1} S/cm**
- **ADDITIONAL INTERESTING FEATURE:**
  $Ba_2In_2O_5$ may absorb water to fill oxygen vacancies → proton conductor



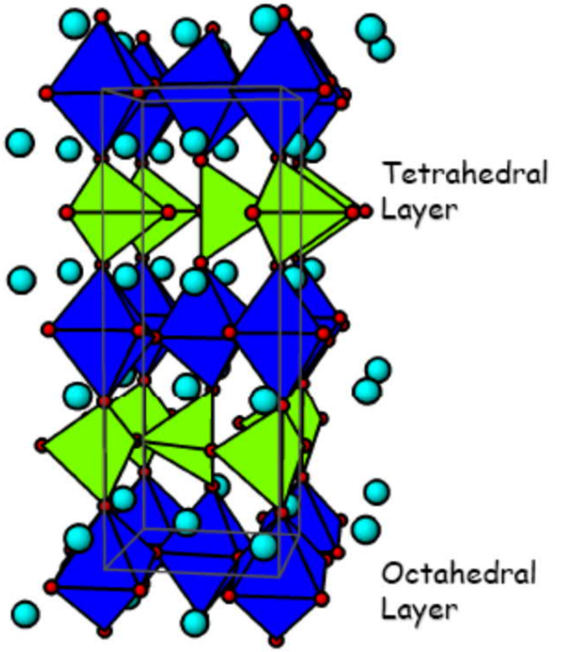


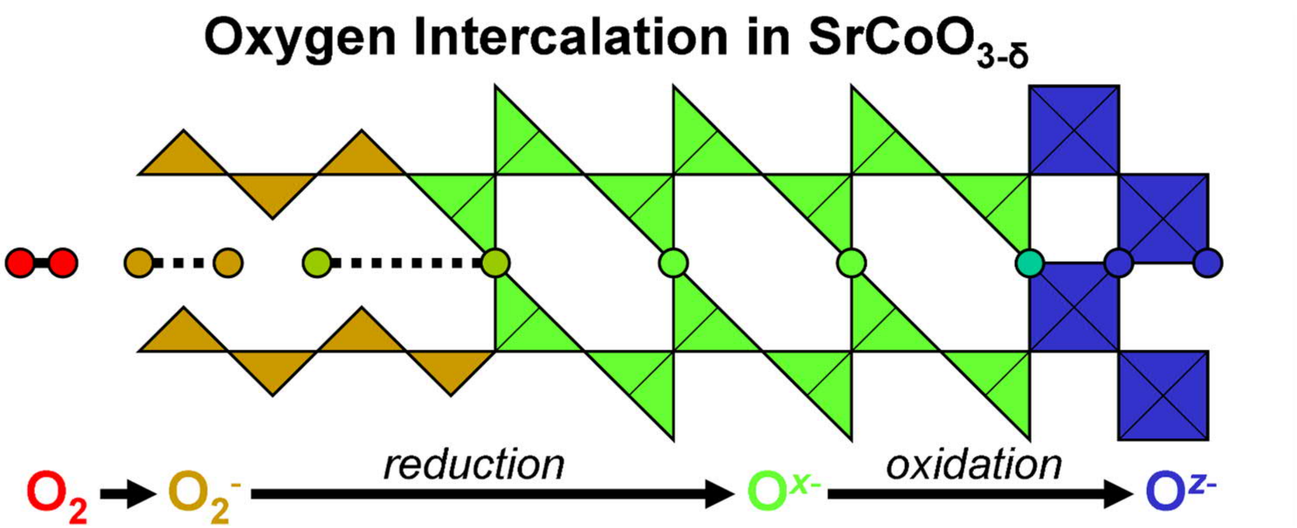
O2
1. gas phase diffusion
O
O
e-
e-
2. adsorpotion
3. ionization
surface reaction
O2-
4. bulk diffusion
O2-
5. ion transfer
O2-

O-*K* and Co-*L* XANES spectroscopy:

Perovskite (brownmillerite) $SrCoO_{3-\delta}$ upon increasing oxygen content from 2.5 to 2.8:

- **O_2 is first absorbed on the surface as O_2^-**
- **then reductively split into O^{x-}**
- **finally in the bulk reoxidized to O^{z-} ($0 < z < x$)**



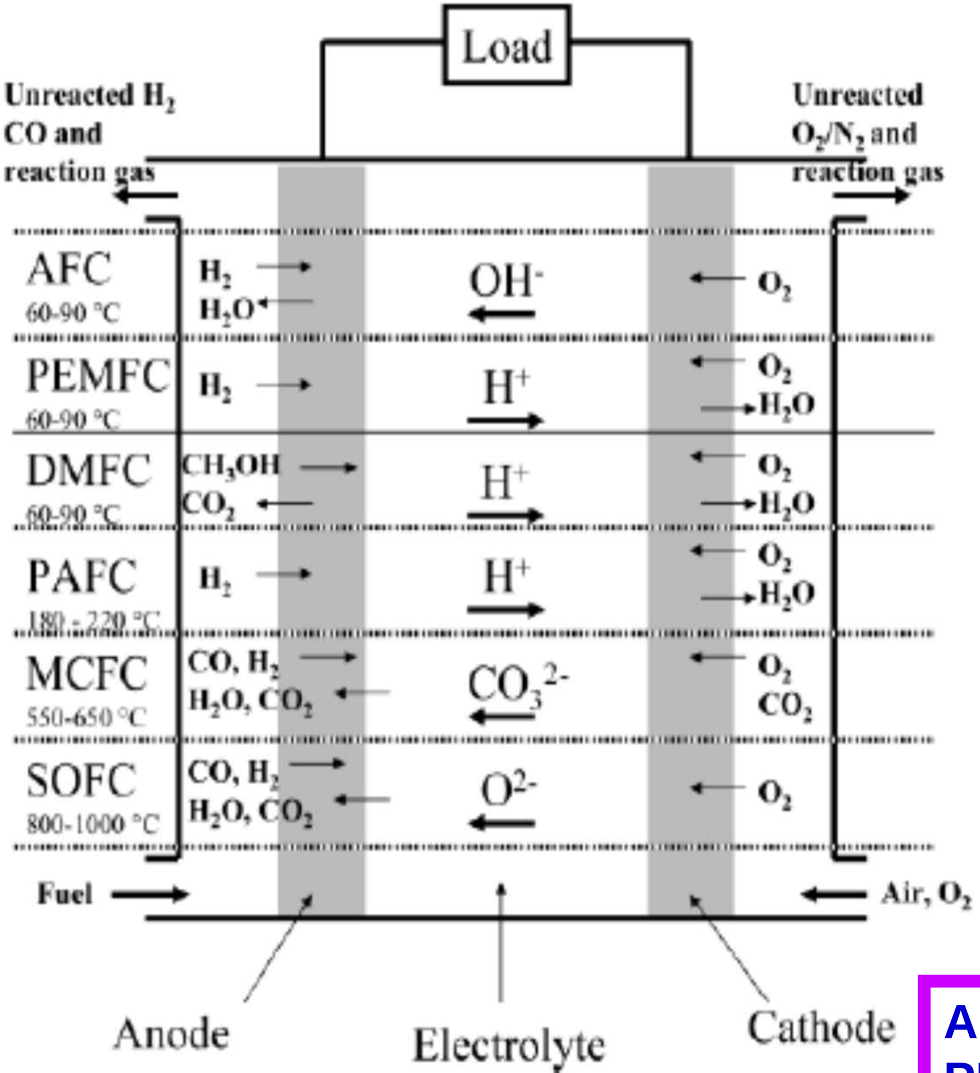


L. Karvonen, M. Valkeapää, R.S. Liu, J.M. Chen, H. Yamauchi & M. Karppinen,
O-*K* and Co-*L* XANES study on oxygen intercalation in perovskite $SrCoO_{3-\delta}$, *Chemistry of Materials* **22**, 70 (2010).

FUEL CELLS

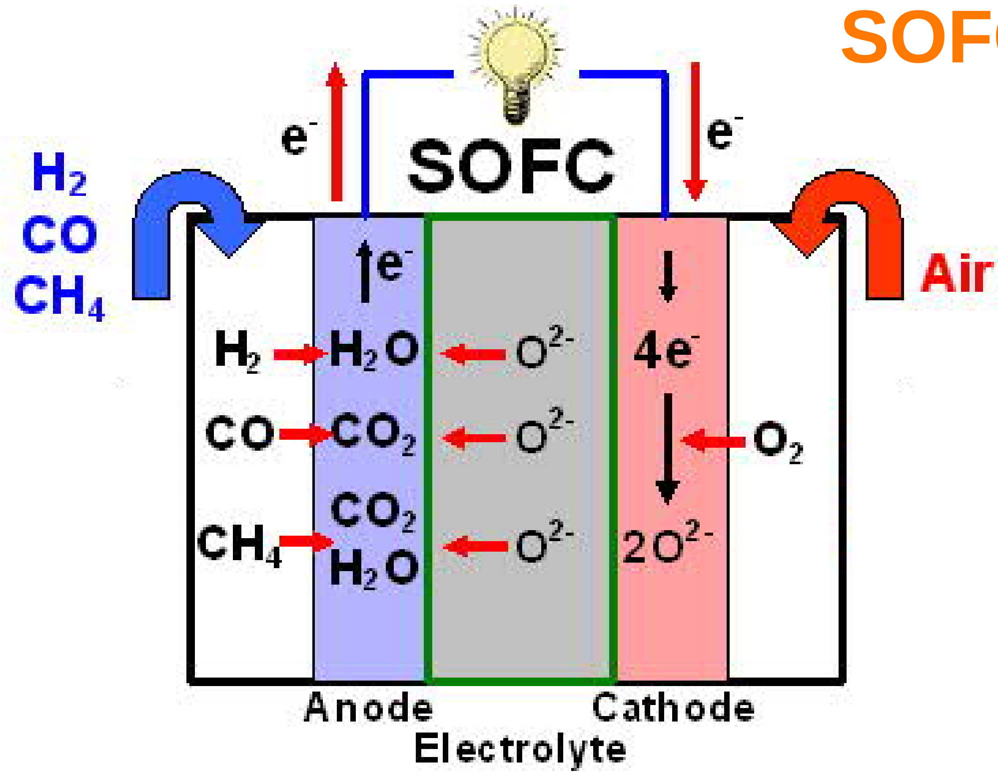
- **ELECTROCHEMICAL DEVICE:** converts the chemical energy from **continuously fed fuel** into electricity through a chemical **reaction with oxygen** or another oxidizing agent
- **POSSIBLE FUELS:** H_2, NH_3, carbon, CO, CH_4, CH_3CH_2OH, propane, butane, natural gas, diesel, Al, Mg, Zn
- Principle of fuel cell: Schönbein 1838
- First practical fuel cell: Bacon 1959
- First applications: Apollo space crafts by NASA in 1970s
- Potential (future) applications: mobile devices, cars, ships, combined heat and electricity production for buildings, …
- **LOW-TEMPERATURE FUEL CELLS:** the main problem is the slowness of oxidation reactions → Pt catalyst → nanostructuring to increase the active Pt surface
- **HIGH-TEMPERATURE FUEL CELLS:** no need for a catalyst, but the thermodynamic efficiency decreases with increasing temperature

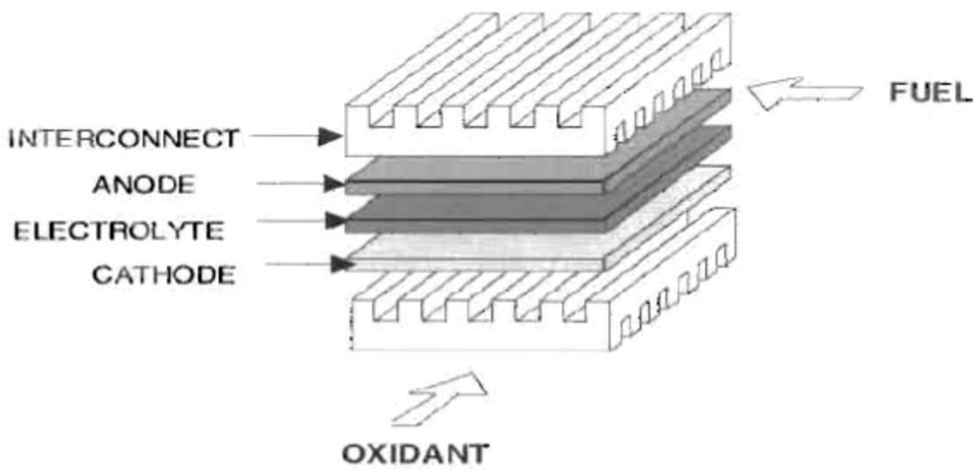
FUEL CELLS

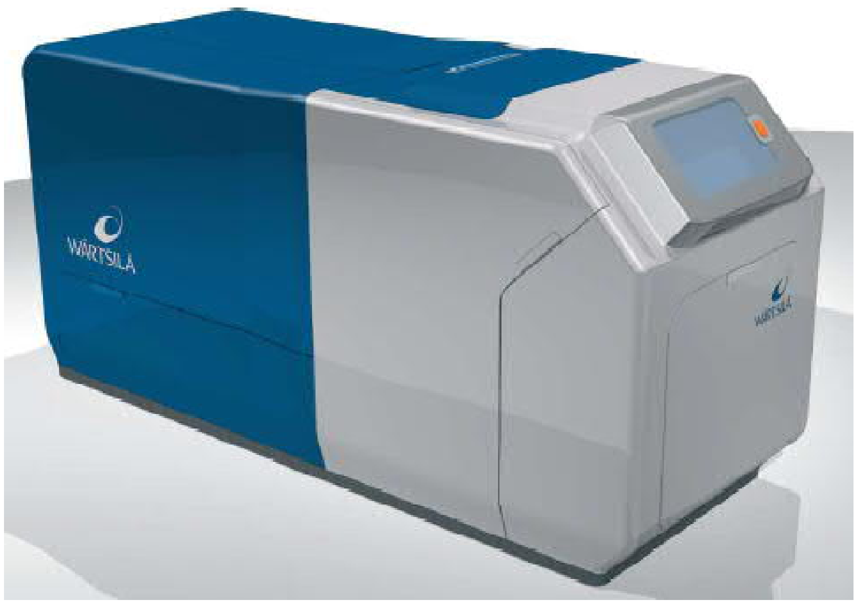


AFC	**Alkaline fuel cell**
PEMFC	**Proton exchange membrane**
DMFC	**Methol direct oxidize fuel cell**
PAFC	**Phosphate fuel cell**
MCFC	**Molten carbonate fuel cell**
SOFC	**Solid-oxide fuel cell**

SOFC: Solid Oxide Fuel Cell







SOFC Material Requirements

ANODE & CATHODE

- **MIEC: Mixed Ionic and Electronic (1 – 100 S/cm) Conductor**
- Chemical & mechanical stability (at 600-900 ^{o}C) under oxidizing conditions for cathode (oxides), under highly reducing conditions for anode (metals); **no coking or sulfur poisoning for anode**
- **Cathode p-type, anode n-type !!!**
- Thermal expansion coefficients to match with the electrolyte
- **Sufficient porosity** to facilitate transport of O_2 gas

ELECTROLYTE

- **High oxide ion conductivity but very low electronic conductivity**
- Stable in both reducing and oxidicing conditions (pO_2: 10^{-20}–1 atm)
- **Free of porosity**

INTERCONNECT (between anode and cathode): **stainless steel or (La,Sr)CrO_3**

- High electronic conductivity and negligible ionic conductivity
- **Free of porosity**
- Stable in both oxidizing and reducing conditions
- Chemical and thermal **compatibility** with other components

- **ELECTROLYTE**
  - oxide-ion conductor & electrical insulator
  - $(Zr,Y)O_2$ (= YSZ; cubic structure & oxygen vacancies)

    (works well only at high operation temperatures)
  - $(La_{0.2}Sr_{0.8})(Ga_{0.3}Mg_{0.7})O_{3-\delta}$ (Ga is expensive)
  - $YBaCo_4O_{7+\delta}$ [M. Karppinen, et al., *Chem. Mater*. 18, 490 (2006)]
- **ANODE**
  - MIEC (mixed ionic & electronic conductor)
  - Ni/YSZ composite

    (works with H_2, but not for C- and S-containing fuels)
  - $(La,Sr)_{0.9}(Cr_{0.5}Mn_{0.5})O_{3-\delta}$ [S.W. Tao & J.T.S. Irvine, *Nature Mater.* 2, 320 (2003)]
  - $Sr_2(Mg,Mn)MoO_{6-\delta}$ [Y.H. Huang, J.B. Goodenough, *et al*., *Science* 312, 254 (2006)]
- **CATHODE**
  - MIEC (mixed ionic & electronic conductor)
  - $(La,Sr)MnO_{3-\delta}$ (reacts with the electrolyte)
  - $(Sr,Ba)(Co,Fe)O_{3-\delta}$ [Z.P. Shao & S.Haile, *Nature* 431, 170 (2004)]

PRESENT ELECTROLYTE: Y-STABILIZED ZIRKONIA (YSZ)

- ZrO_2 – 8% Y_2O_3: cubic fluorite structure
- Y^{3+}-for-Zr^{4+} substitution creates oxygen vacancies
- **Electronic conductivity low enough**
- Good mechanical properties & relatively low price
- **PROBLEM: oxide-ion conductivity somewhat low**
  (could be improved by *e.g*. replacing Y with Sc, but Sc very rare/expensive)

NEW ELECTROLYTE CANDIDATE: $(La,Sr)(Ga,Mg)O_{3-\delta}$

- Perovskite structure
- Sr^{2+}-for-La^{3+} & Mg^{2+}-for-Ga^{3+} → **oxygen vacancies** → **oxide-ion conductivity**
- **Electronic conductivity low enough**
- **PROBLEMS:**
  - Decomposes at high temperatures & reducing conditions → operation < 800 °C
  - Mechanical properties not as good as for YSZ
  - Reacts with some electrode materials → buffer layers ?
  - Ga is expensive → Al-for-Ga substitution ?

PRESENT CATHODE: (La,Sr)MnO$_3$

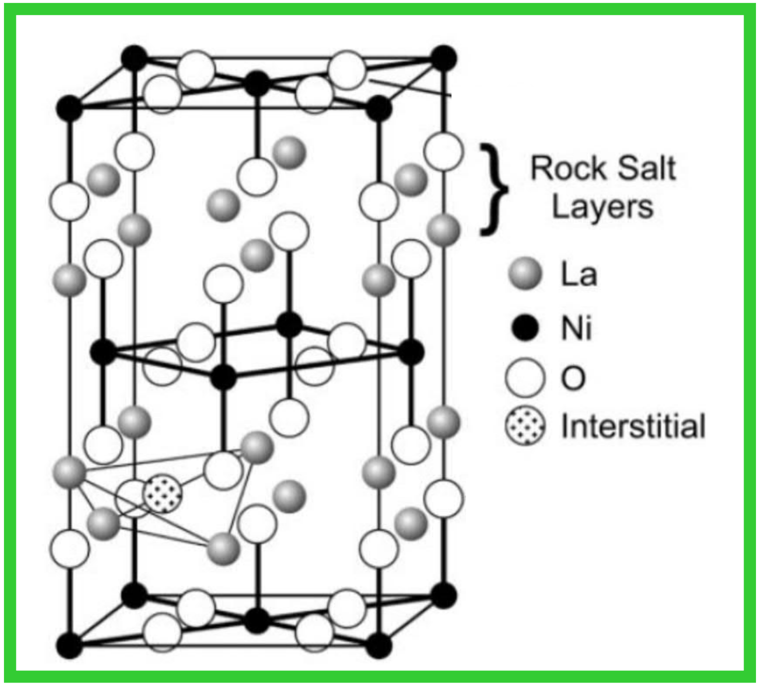
- Perovskite structure
- Sr^{2+}-for-La^{3+} substitution: **Mn^{3+} → Mn^{4+}**
  → **Good electronic conductivity** (σ = 500 S/cm @ 800 °C)
- Stable under highly oxidizing conditions & Low price
- **PROBLEM: too low ionic conductivity**
  → mixing with electrolyte for a composite

NEW CATHODE CANDIDATE: (La,Sr)(Co,Fe)O$_{3-\delta}$

- Perovskite structure
- (La,Sr)**Co**O$_{3-\delta}$
  - Sr^{2+}-for-La^{3+} substitution: **Co^{3+} → Co^{4+} & Oxygen vacancies**
  - **Very good MIEC**
  - **PROBLEM: Co is expensive**
- (La,Sr)**Fe**O$_{3-\delta}$
  - Fe much cheaper than Co
  - Better thermal expansion characteristics
  - But lower electrical conductivity

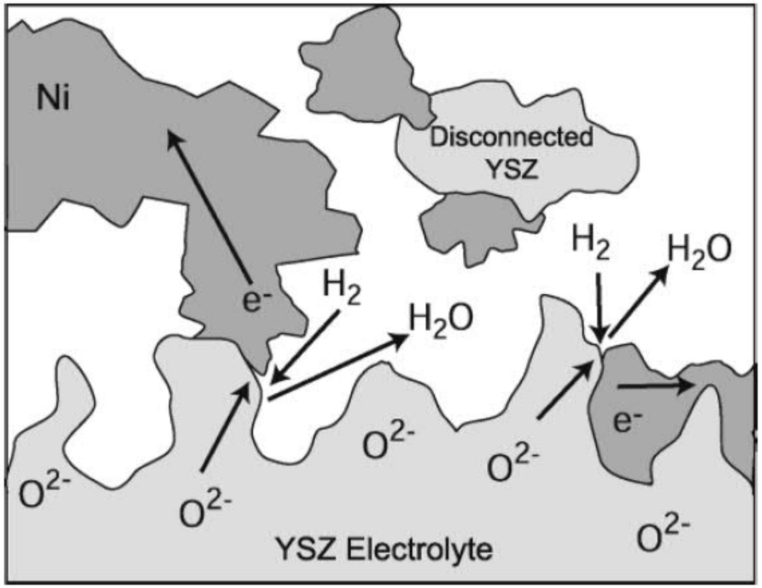
NEW CATHODE CANDIDATE: La_2NiO_4

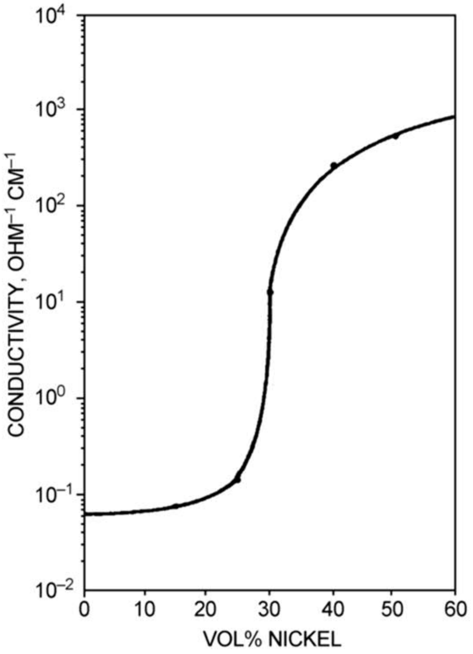
- Ruddlesden-Popper (RP) structure
- Excellent electrical conductivity: semiconductor-metal transition around 400°C
- Very good ionic conductivity: **interstitial oxygen !**
- Reacts with YSZ; long term stability ?



PRESENT ANODE: NICKEL

- Reducing conditions → metals → Ni best
- High electronic conductivity (~10^5 S/cm)
- Low price
- **No oxide-ion conductivity** → **mixing with electrolyte for a composite (cermet)**
- Works perfectly with H_2 as a fuel
- **Poisoning when the fuel contains C or S**





NEW ANODE CANDIDATE: $(La,Sr)CrO_3$

- $LaCrO_3$: too low conductivity
- Sr^{2+}-for-La^{3+}: $Cr^{3+} \rightarrow Cr^{4+}$, increased electrical conductivity (but *p*-type, **not good!)**
- Decent sulfur tolerance

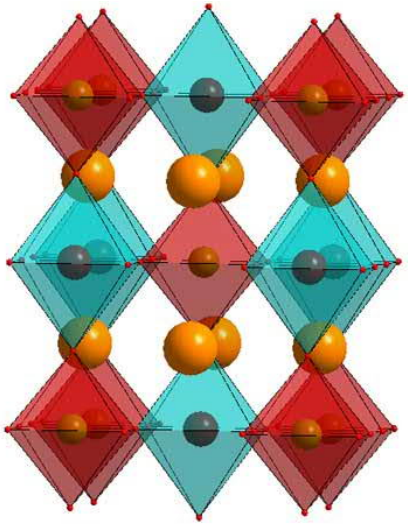
NEW ANODE CANDIDATE: Sr_2MgMoO_6

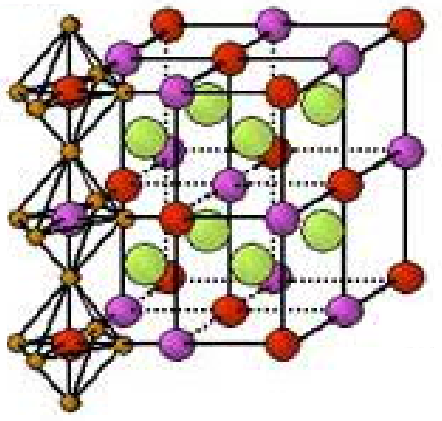
- *B*-site ordered double perovskite, **n-type**
- Decent electronic conductivity (Mo^{6+} → Mo^{5+}; σ = 1-10 S/cm)
- Stable in both reducing and oxidizing conditions
- **Stable with C and S containing fuels**
- Substitution: La^{3+} → Sr^{2+}
  - Mo^{6+} → Mo^{5+}
  - Works perfectly with hydrocarbons

- **VERY PROMISING !**

Y.H. Huang, **J.B. Goodenough**, *et al*., *Science* **312**, 254 (2006).

$A_2B'B''O_6$



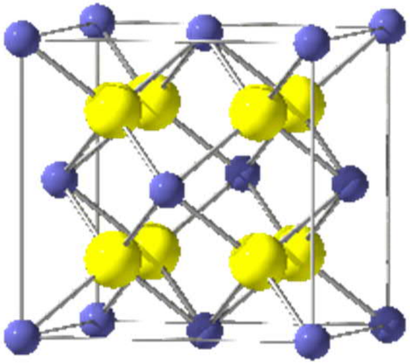


Sr_2FeMoO_{6-w}

Sr_2MgMoO_{6-w}

Y.H. Huang, R.I. Dass, Z.L. Xing & J.B. Goodenough,
Double perovskites as anode materials for solid-oxide fuel cells,
Science **312**, 254 (2006).

OXYGEN STORAGE



COMMERCIAL CeO_2

- $CeO_{2-\delta}$: $Ce^{III/IV}$
- $(Ce,M)O_{2-\delta}$: *M* = Zr, Ti, Y, Bi, etc. (commercial)
- OSC ≈ 1500 μmol-O / g_{katal} (500 °C)

  [Y. Nagai *et al.*, *Catalysis Today* 74, 225 (2002)]
- good oxide-ion conductivity when T >500 °C
  → exhaust catalyst, SOFC electrolyte

NEW OXYGEN-STORAGE MATERIAL !!

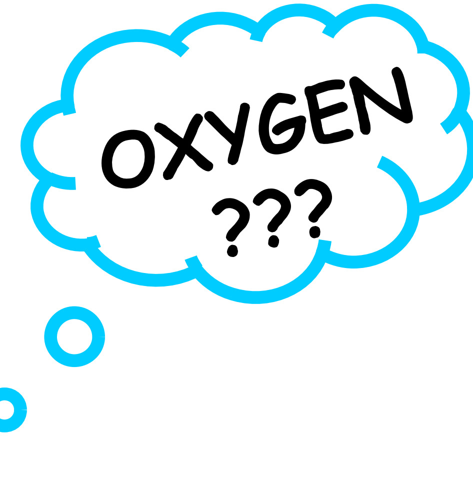
- $YBaCo_4O_{7+\delta}$: $Co^{II/III}$ ($0 < \delta < 1.5$)
- OSC ≈ 2700 μmol-O / g_{katal} (200 ~ 350 °C)

  [M. Karppinen *et al*., *Chem. Mater*. **18**, 490 (2006);
  Int. Patent Appl. PCT/JP2006313436, filed June 6, 2006]

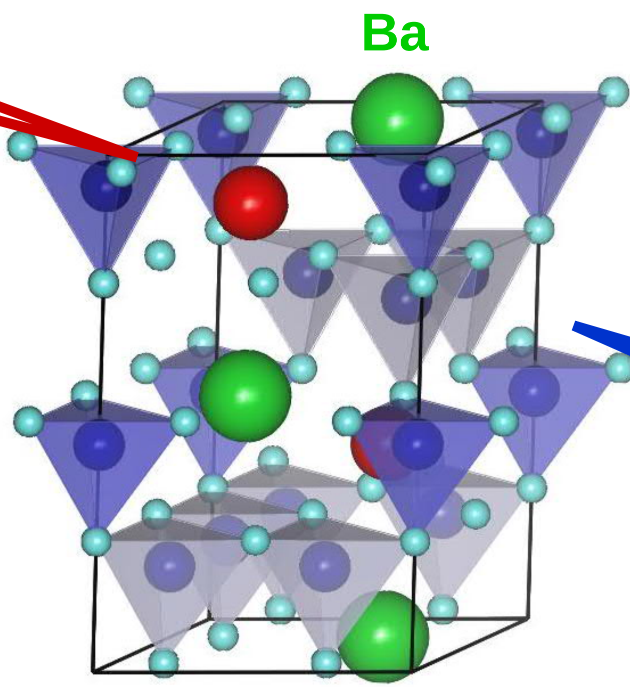
OSC (oxygen-storage capacity): **μmol O / g**

$YBaCo_4O_7$

- **Compound discovered in 2002 in Sweden**
  [M. Valldor & M. Andersson, *Solid State Sci.* 4, 923 (2002).]
- **Investigated for thermoelectric properties (layered Co oxide)**
- **Investigated for magnetic properties (frustrated Kagome-lattice)**



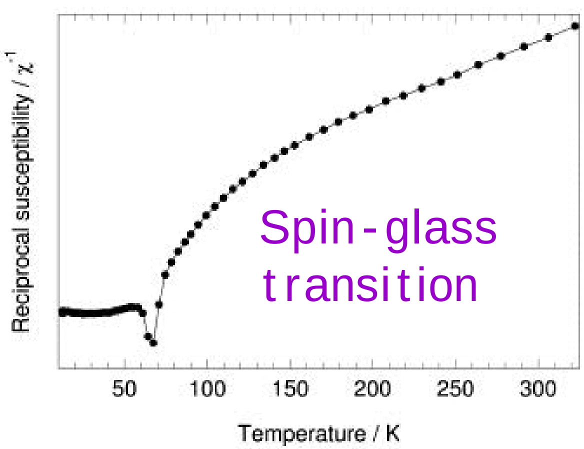
Y: Dy ~ Lu, Ca, In



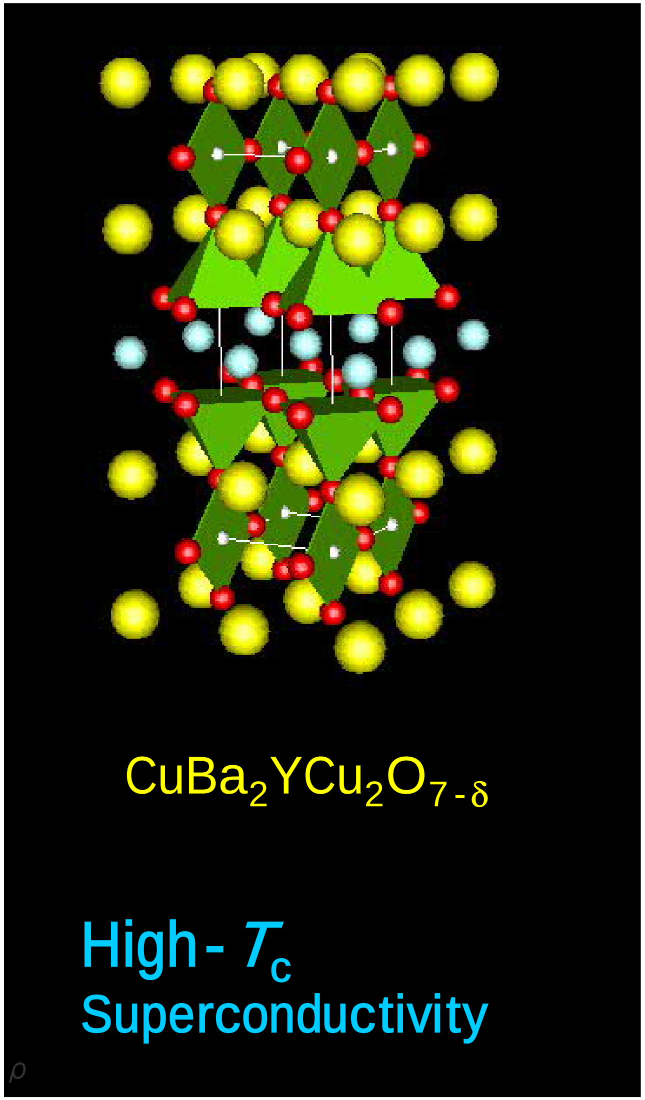
Triangular Layer (Co^{III})

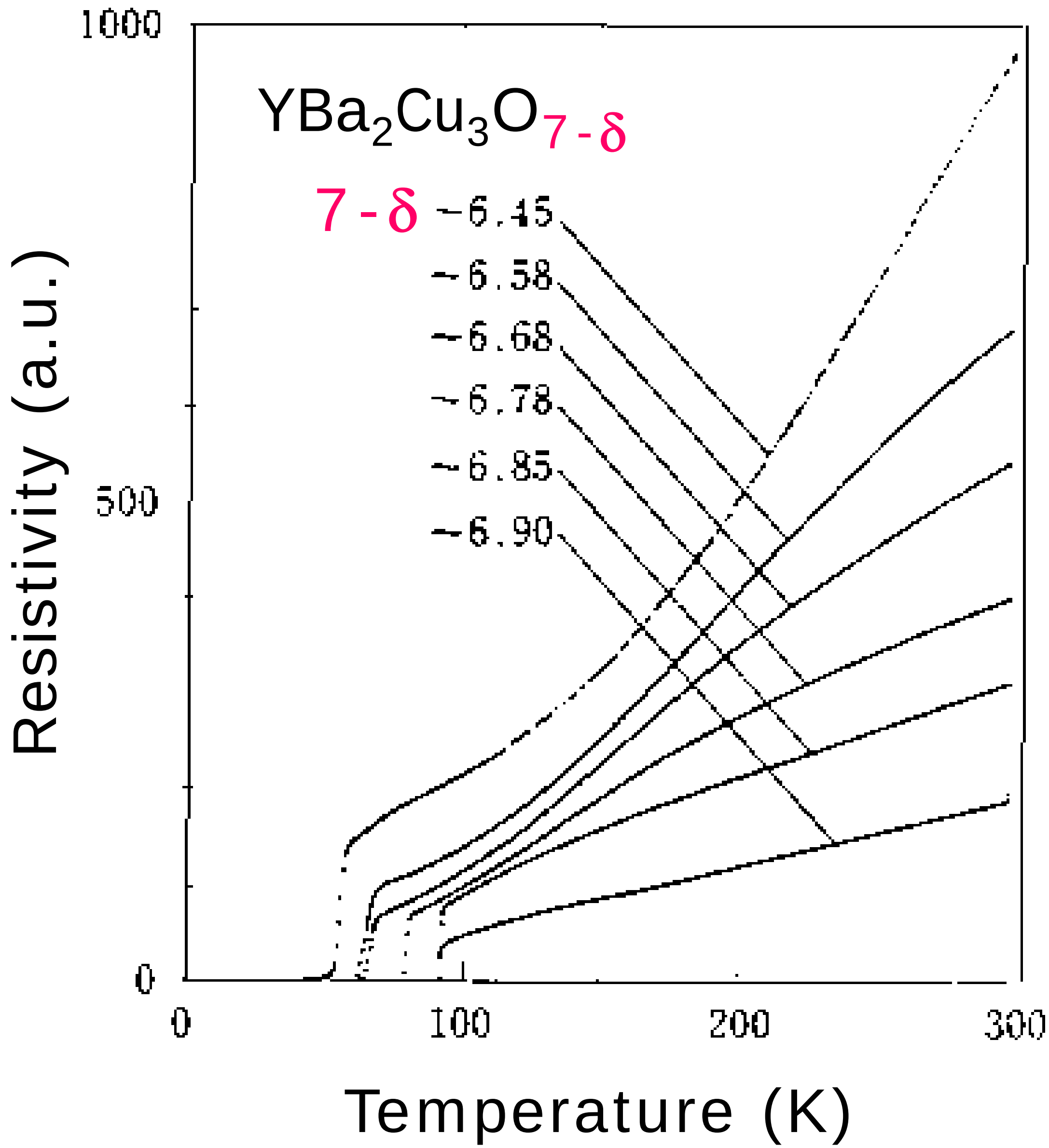
Kagomé Layer (Co^{II})

Co: Al, Zn, Fe



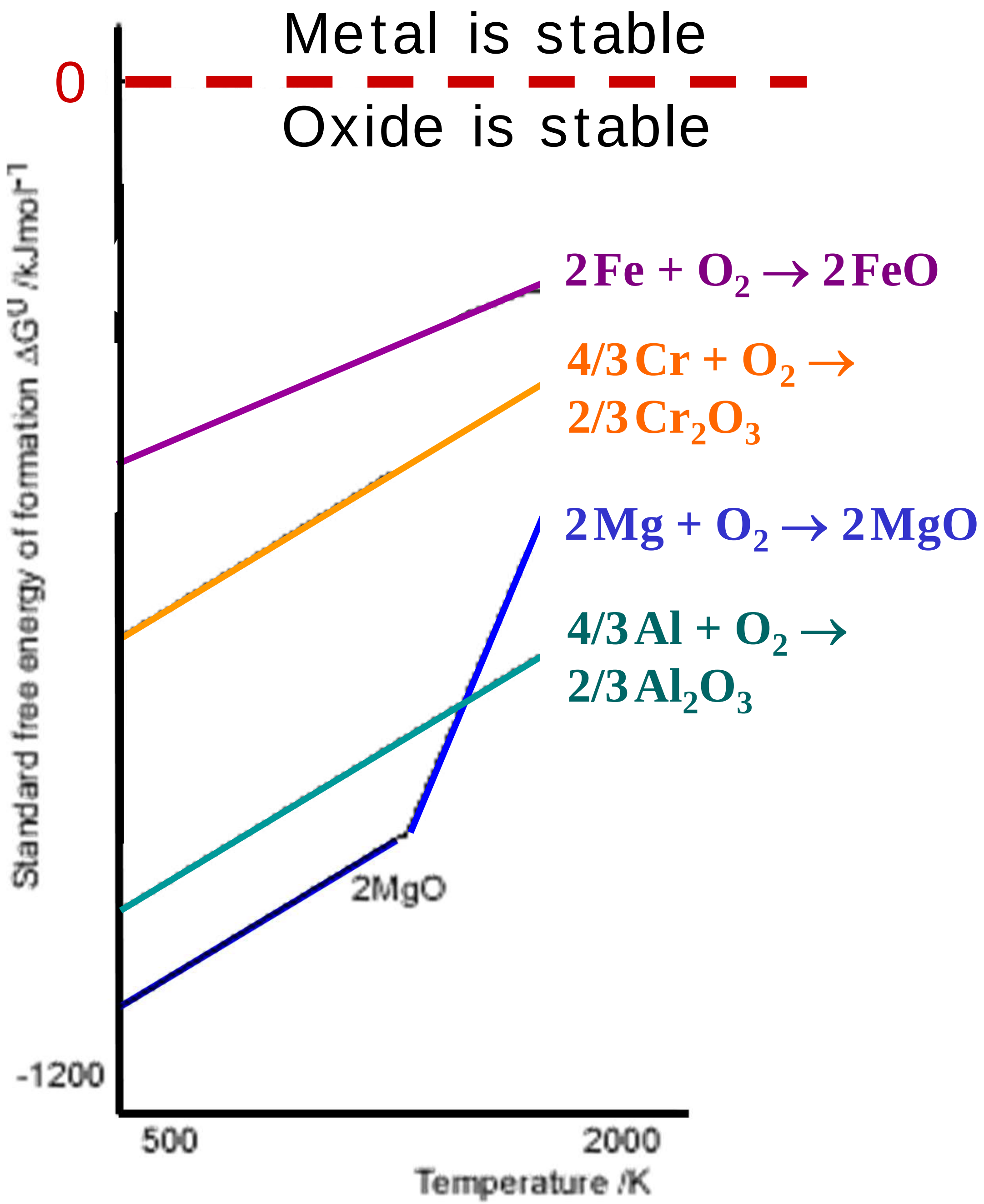
SUPERCONDUCTIVITY depends on OXYGEN CONTENT

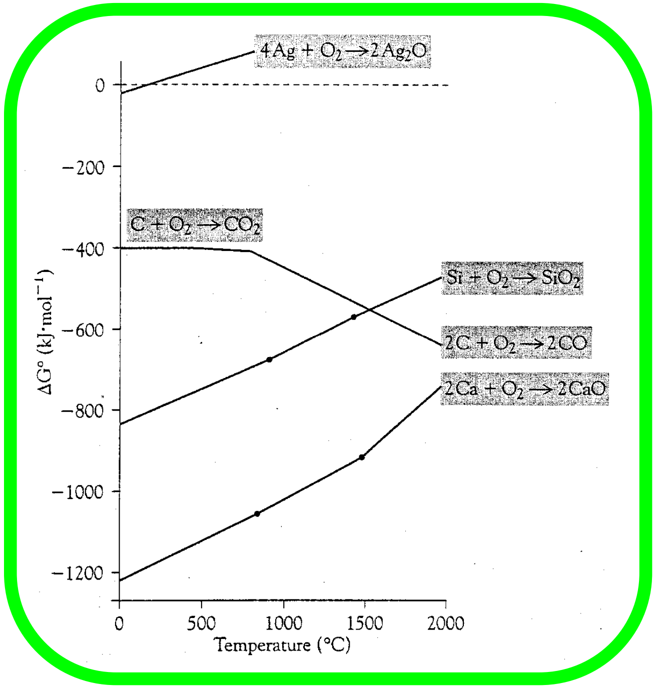




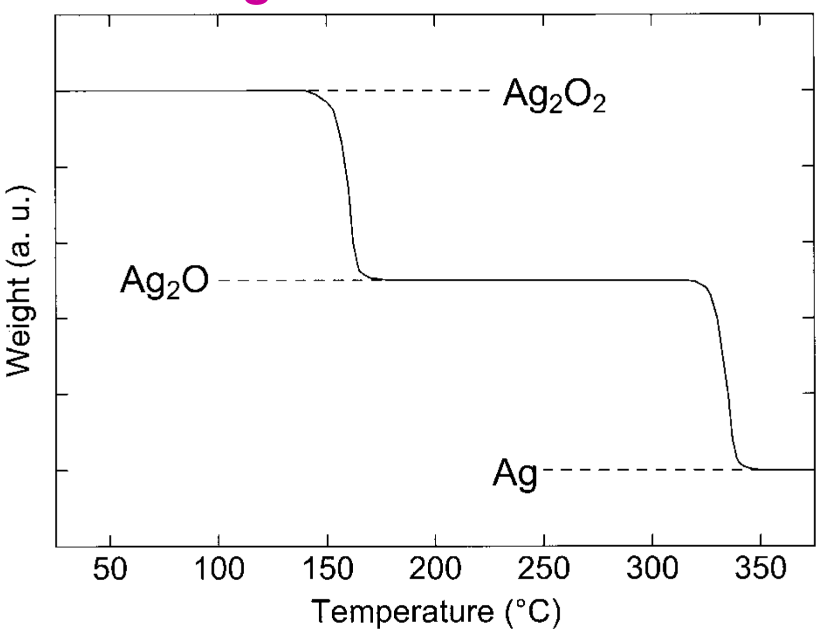
Ellingham diagram

- (Gibb's) free energy of formation versus temperature for metal oxides
- Temperature at which a metal oxide is spontaneously reduced to a metal

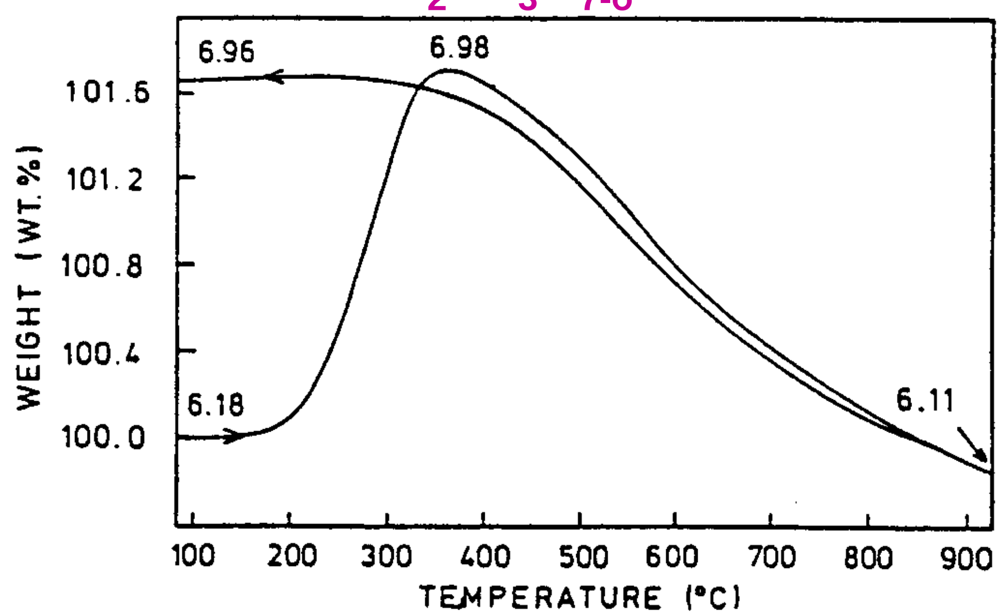




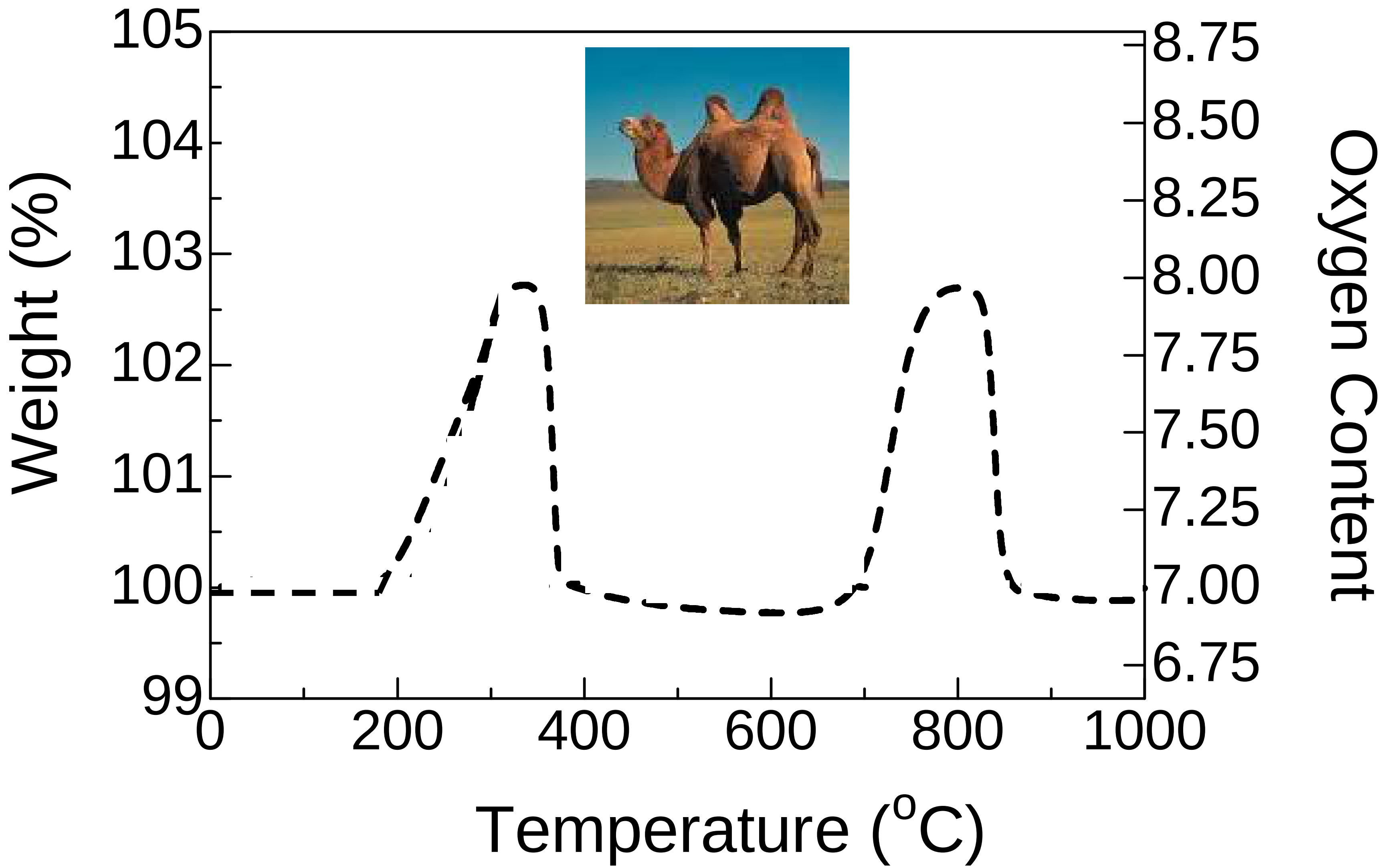
TG of AgO in air

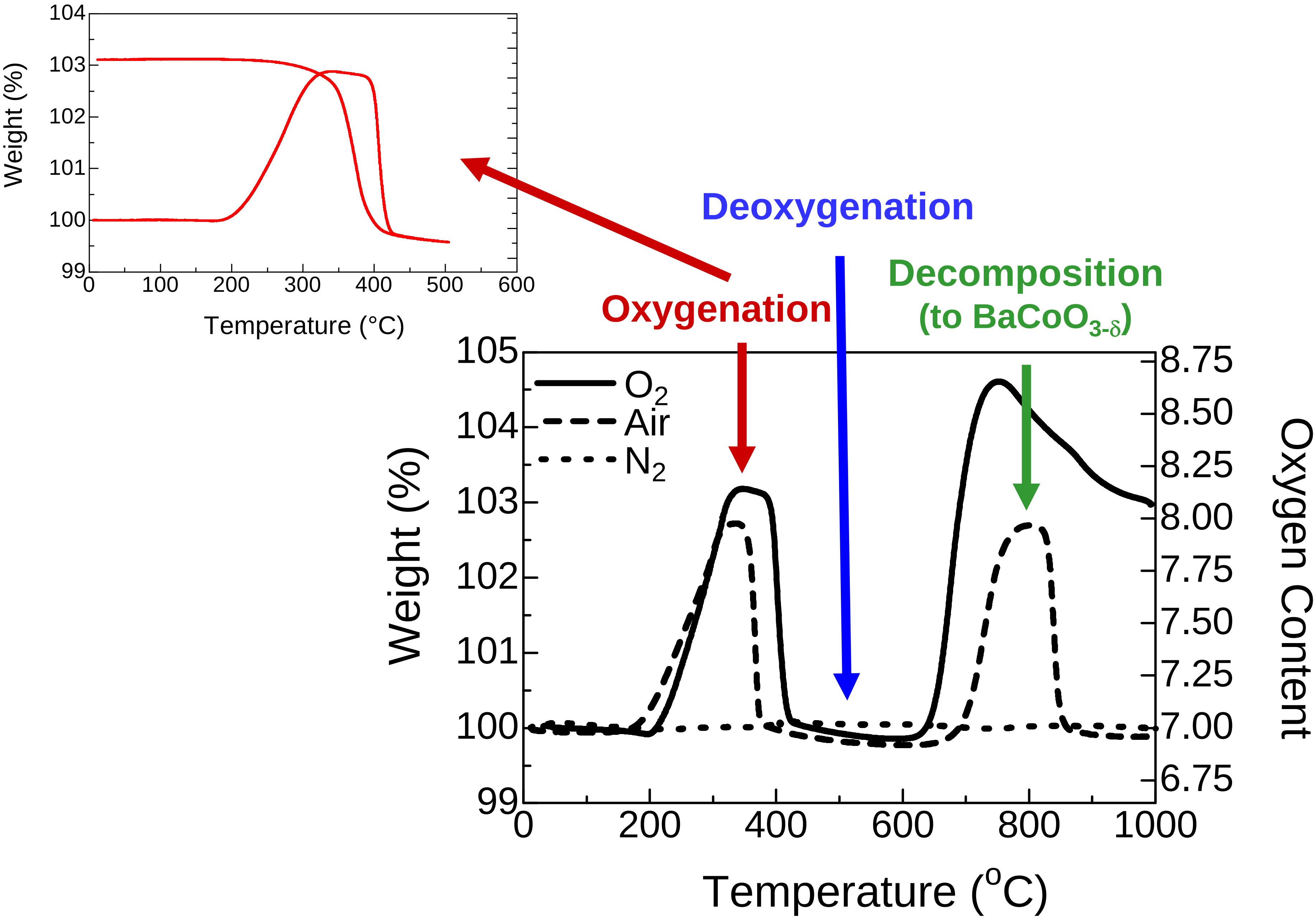


TG of $YBa_2Cu_3O_{7-\delta}$ in air



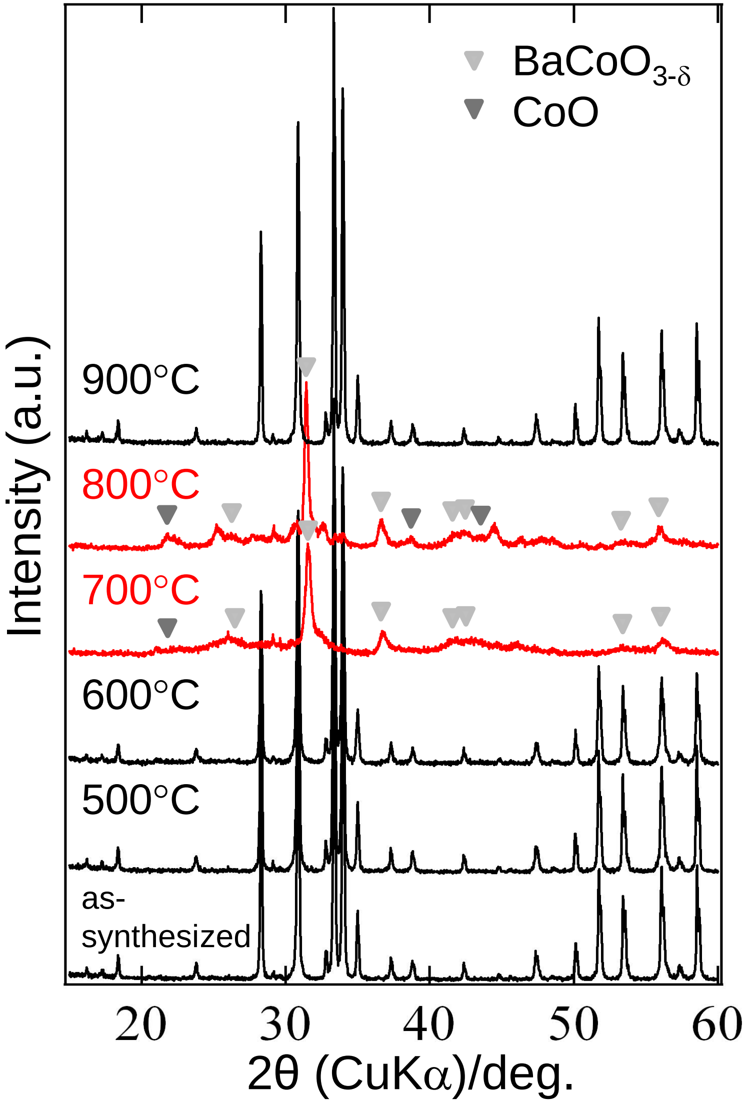
$YBaCo_4O_7$: heating in air in a thermobalance

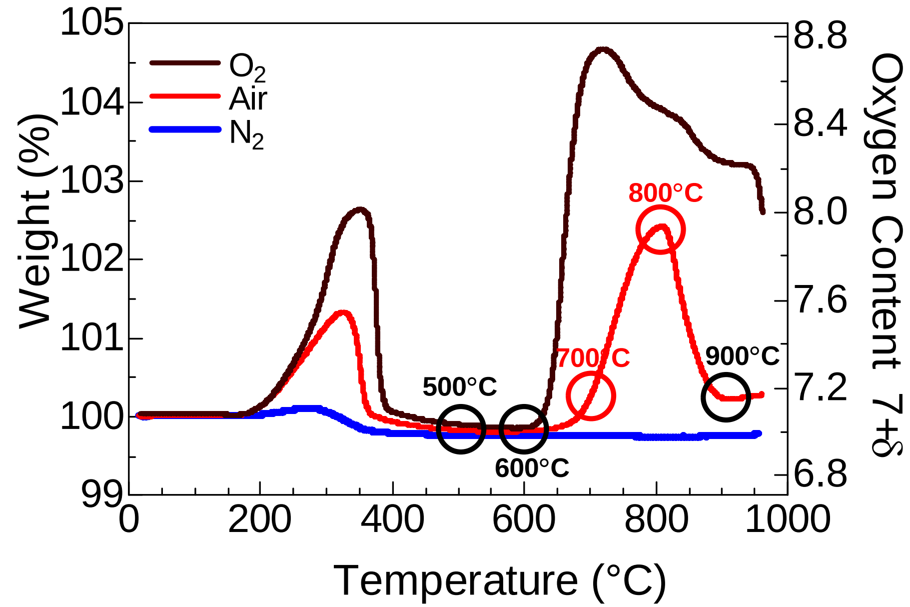




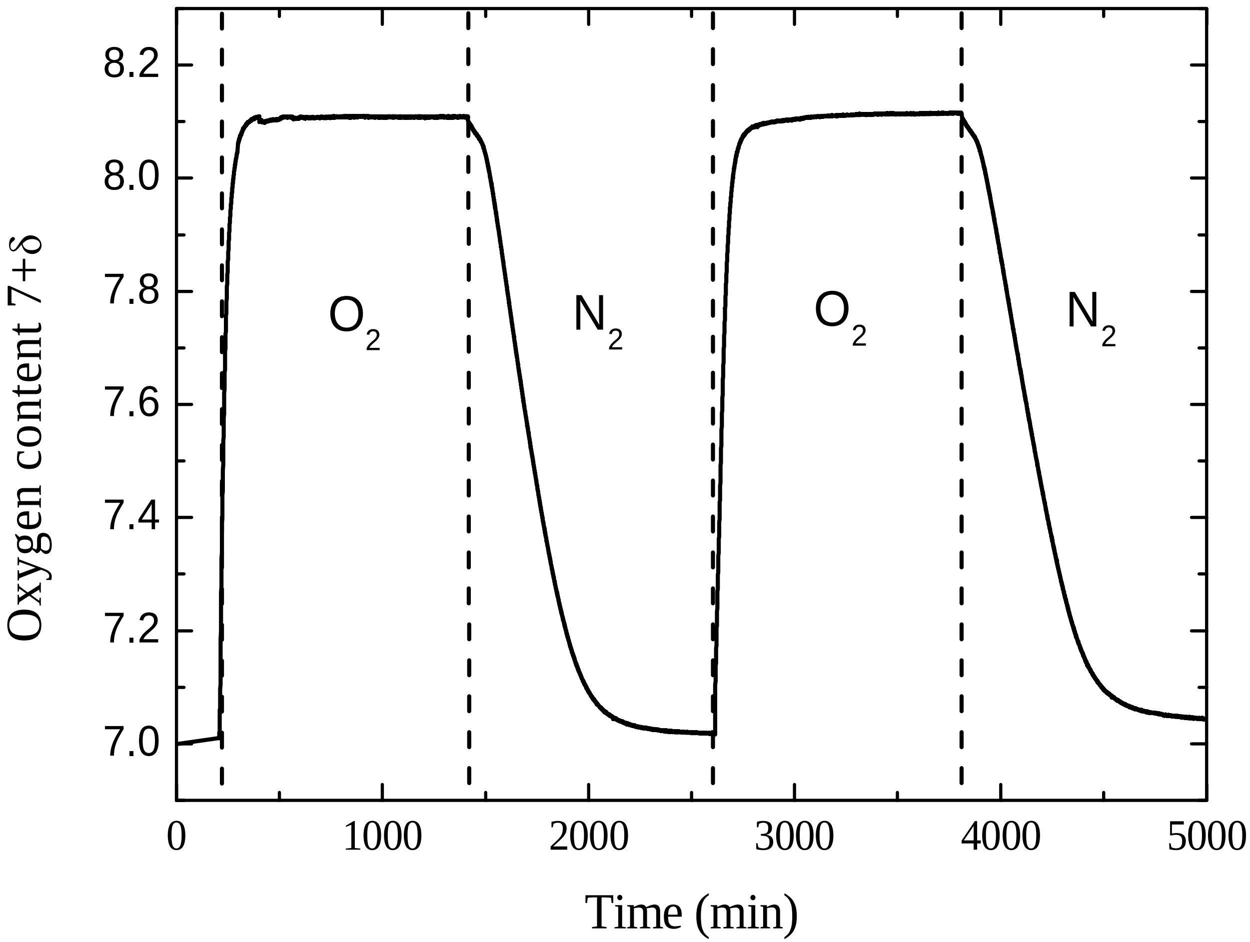
M. Karppinen, H. Yamauchi, S. Otani, T. Fujita, T. Motohashi, Y.-H. Huang, M. Valkeapää & H. Fjellvåg, *Chem. Mater*. **18**, 490 (2006).

Decomposition of $YBaCo_4O_{7+\delta}$ at high temperatures

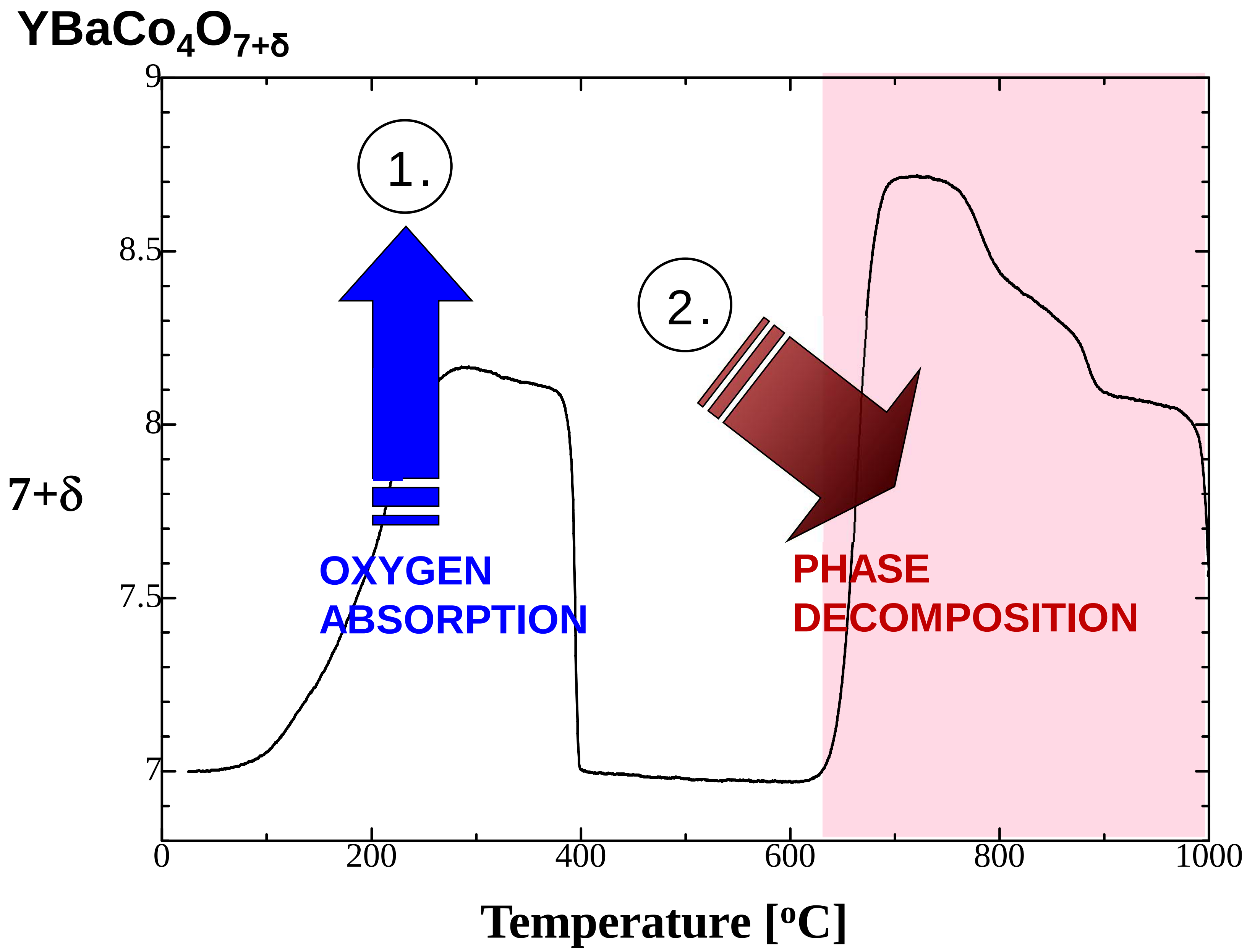




Temperature: 300 ^oC



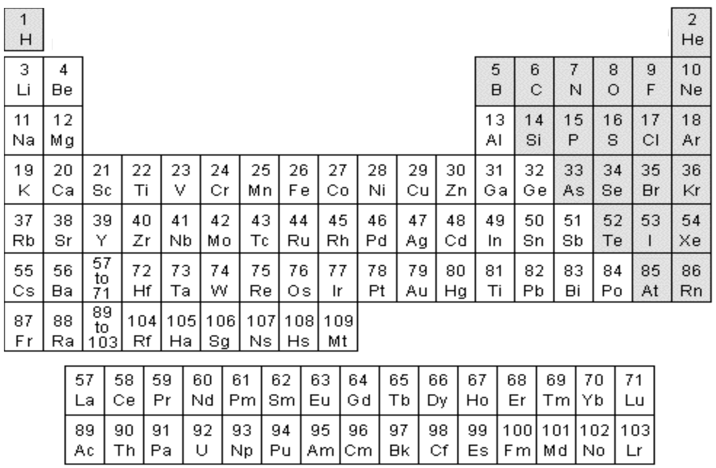
Further Improvements

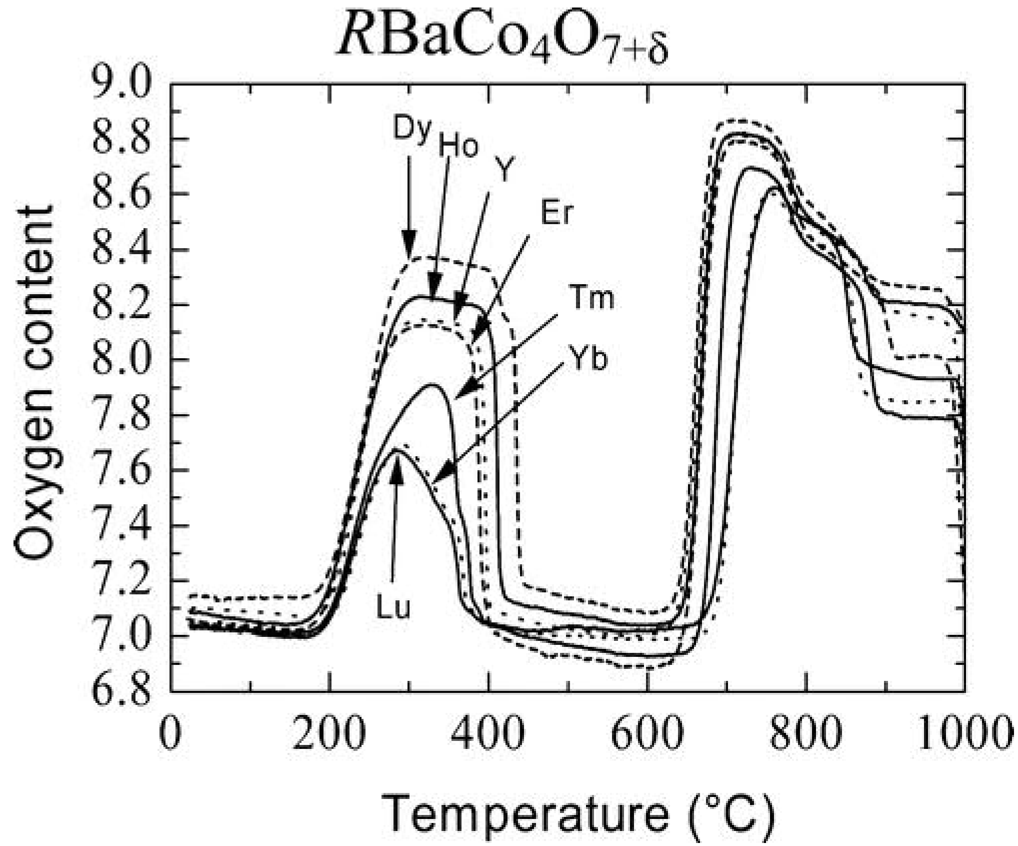


RBaCo$_4$O$_{7+\delta}$

***r*(*R*^{III})** **ionic radiys decreases**

→ decomposition temperature increases (but OSC decreases)



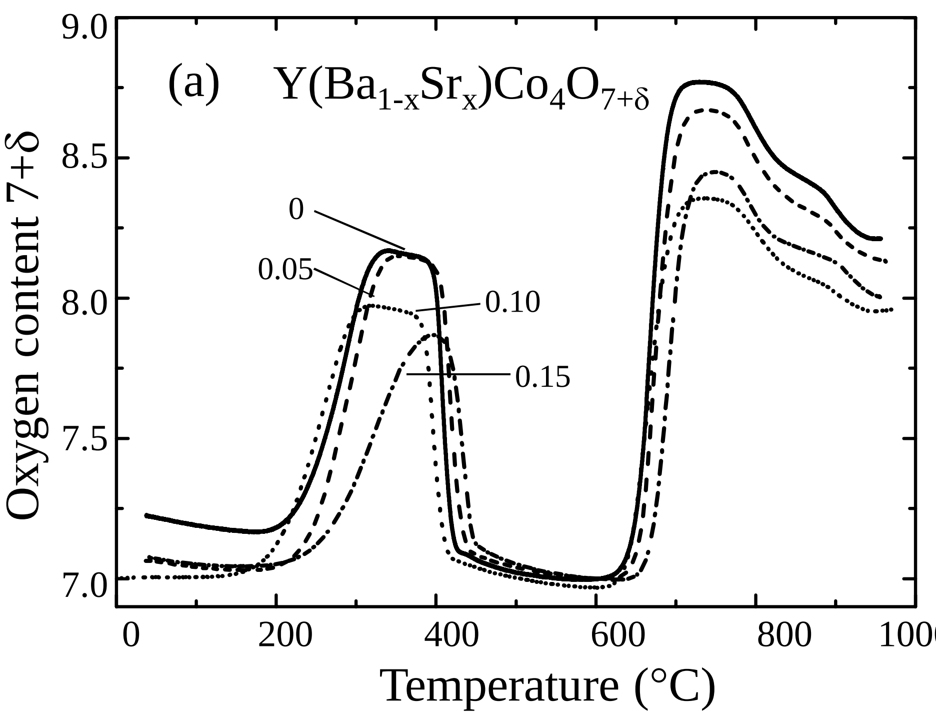


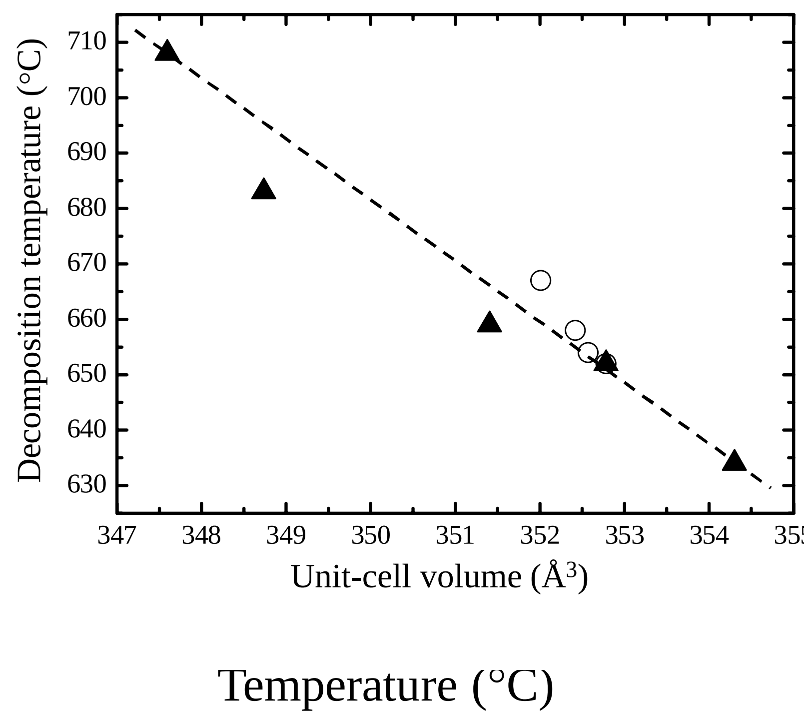
S. Kadota, M. Karppinen, T. Motohashi & H. Yamauchi, *Chem. Mater*. **20**, 6378 (2008).

Y(Ba$_{1-x}$Sr$_x$)Co$_4$O$_{7+\delta}$

***r*(Ba,Sr) ionic radius decreases**

→ decomposition temperature increases (but OSC decreases)

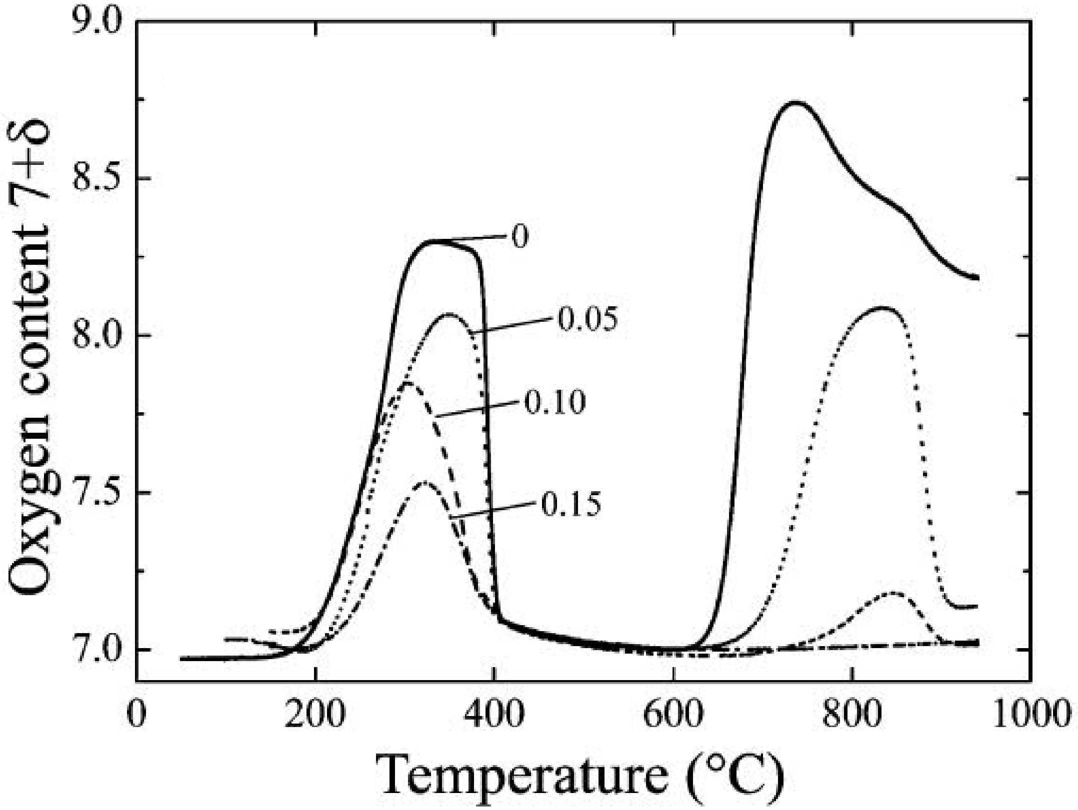




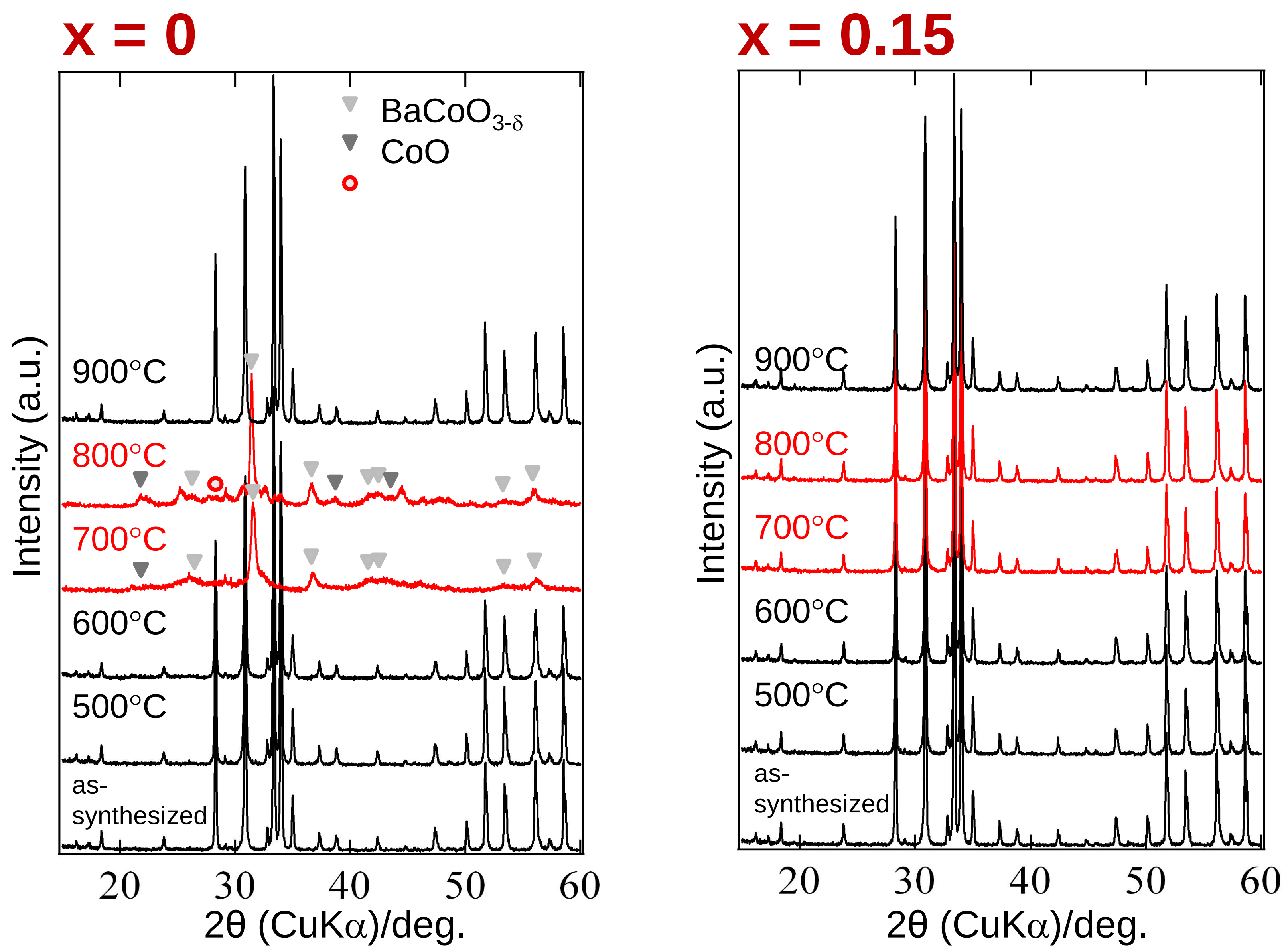
O. Parkkima, H. Yamauchi & M. Karppinen, *Chem. Mater.* **25**, 599 (2013).

$YBa(Co_{1-x}Al_{x/2}Ga_{x/2})_4O_{7+\delta}$

Al-substitution: x < 0.10
Ga-substitution: x < 0.25



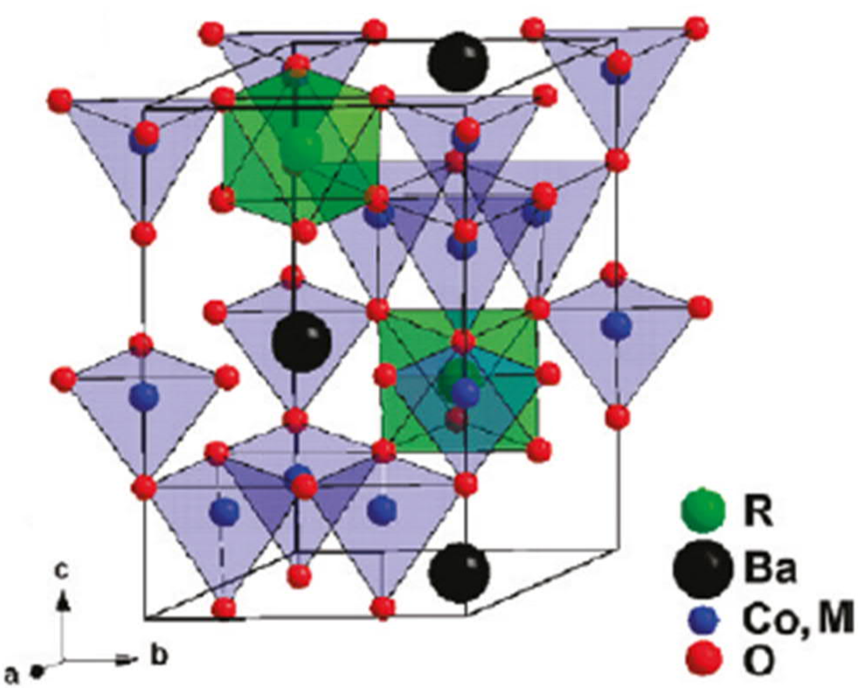
O. Parkkima, H. Yamauchi & M. Karppinen, *Chem. Mater.* **25**, 599 (2013).



$YBa(Co_{1-x}Al_{x/2}Ga_{x/2})_4O_{7+\delta}$

$YBaCo_4O_7$: NEW TYPE OF ELECTROLYTE ?

- Excellent oxide-ion conductivity (σ_i > 30 S/cm)
- Stability issues: Co → Zn, Al, Ga etc.
- Reactivity with other SOFC materials ???



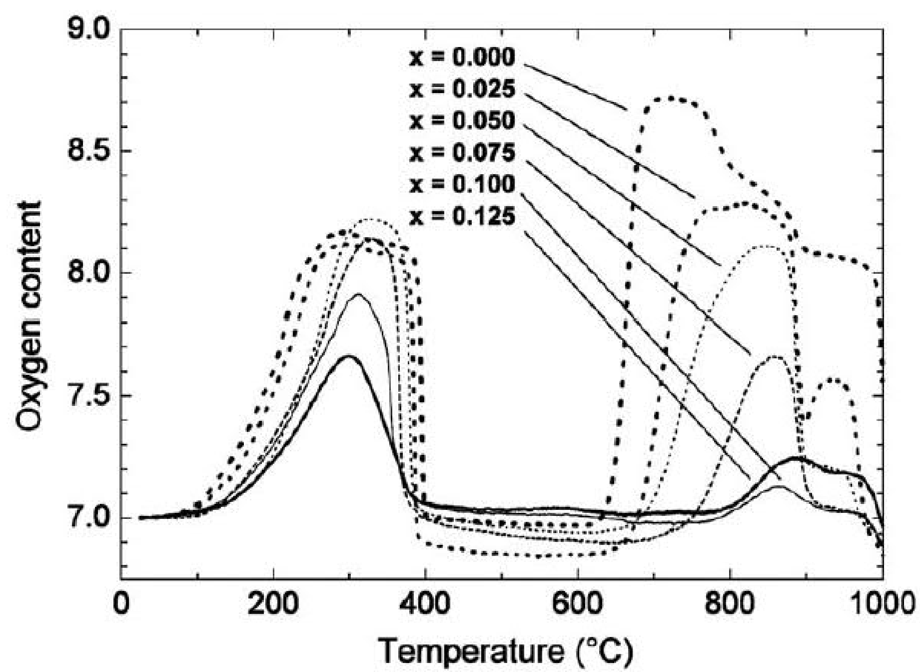


Fig. 3. TG curves for the $YBa(Co_{1-x}Al_x)_4O_{7+\delta}$ samples ($0.00 \leq x \leq 0.125$; $\delta \approx 0$) recorded upon heating the sample material in an O_2 gas flow.

$YBaCo_4O_{7+\delta}$ has already been investigated for:

- H_2/O_2 separation after photocatalytic water splitting (Mitsubishi Chemical Corp.)
- sorbent material for oxygen-enriched CO_2 stream production
- oxygen-separating membrane material
- cathode material for solid oxide fuel cell
- catalyst in epoxidation reaction requiring active oxygen species

[O. Parkkima, A. Silvestre-Albero, J. Silvestre-Albero & M. Karppinen, Superior performance of oxygen-nonstoichiometric $YBaCo_4O_{7+\delta}$ as a catalyst in H_2O_2 oxidation of cyclohexene, *Catal. Lett.* ***145***, 576 (2015)]